

Quantum Numbers

- _____
- Used to _____ an _____ in an _____

n

- _____
- Represents _____ energy level of _____
- _____ # of _____ in an _____

Example: What is the maximum number of electrons that can be in the _____ main energy level?

l

- The _____
- Describes the _____ within an _____
- _____ of orbital _____ possible in _____

Orbital Shapes

designated _____

- level 1: _____
- level 2: _____
- level 3: _____
- level 4: _____

How many electrons can each sublevel hold?

$$s = 1 \text{ orbital} \times 2 \text{ e}^-/\text{orbital} = \text{_____} \text{e}^-$$

$$p = 3 \text{ orbitals} \times 2 \text{ e}^-/\text{orbital} = \text{_____} \text{e}^-$$

$$d = 5 \text{ orbitals} \times 2 \text{ e}^-/\text{orbital} = \text{_____} \text{e}^-$$

$$f = 7 \text{ orbitals} \times 2 \text{ e}^-/\text{orbital} = \text{_____} \text{e}^-$$

Gpb Chemistry Note Taking Guide 1203 Answers

A. Jones



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