

Chemistry Unit 8 Worksheet 4

Samples of Every Kind of Problem

Let us suppose that our "paper" writer is a chemist who has to write out the products of the reaction. Show the given data and find the mass. The mass is underlined (underlined) given.

1. Chlorine gas reacts with a solution of potassium hydroxide, KOH , to form three different products: potassium chloride, KCl , and a potassium compound gas.

Reactants:	Cl_2	KOH
Products:	KCl	K_2O
Masses:	<u>10.0 g</u>	<u>10.0 g</u>
Mass:	10.0	10.0

Reaction:
Product:

2. A sample of a displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. The mass of the magnesium metal is 10.0 g and the mass of the hydrogen gas is 0.5 g.

Reactants:	Mg	HCl
Products:	MgCl_2	H_2
Masses:	<u>10.0 g</u>	<u>0.5 g</u>
Mass:	10.0	0.5

Reaction:
Product:

3. Chlorine, Cl_2 , reacts with a solution of potassium hydroxide, KOH , to form three different products: potassium chloride, KCl , and a potassium compound gas. The mass of the chlorine gas is 10.0 g and the mass of the potassium compound gas is 0.5 g.

Reactants:	Cl_2	KOH
Products:	KCl	K_2O
Masses:	<u>10.0 g</u>	<u>0.5 g</u>
Mass:	10.0	0.5

Reaction:
Product:

4. Chlorine gas reacts with a solution of potassium hydroxide, KOH , to form three different products: potassium chloride, KCl , and a potassium compound gas. The mass of the chlorine gas is 10.0 g and the mass of the potassium compound gas is 0.5 g.

Reactants:	Cl_2	KOH
Products:	KCl	K_2O
Masses:	<u>10.0 g</u>	<u>0.5 g</u>
Mass:	10.0	0.5

Reaction:
Product:

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