For the reverse reaction? <u>(ICC/EaSed</u> The Ea/r)
q) What effect did the catalyst have on the ΔH of the forward reaction? ho effect
The reverse reaction? no effect
12. Name four instances in which catalysts are used in industry or everyday life and tell which catalysts are used. V20s Used in making sulphuric acid/step 2502 + 02 2503
11 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1
mattase breaks down mattose to simple sugar (glucose)
nickel used to hydrogenate unsaturated fats
Pt, Pd, Rh used in catalytic converters to breakdown NO2 in exhaus
 Describe what happens to the kinetic energy, potential energy and the total energy of reactant molecules as they approach each other.
KE decreases PE increases TE remains constant.
I.d. Combain subura bassas materiales access for a constant to the
 Explain why a lower activation energy for a reaction leads to a greater reaction rate at a given temperature.
molecular collisions of require as high an energy to be surressful
molecular collisions of require as high an energy to be successful so of a particular temp. a greater fraction of the molecules will have enough energy. A small piece of this react with 2.0 M HCI to produce 12.0 mL of Hz was in 30.0 seconds.
15. A small piece of zigic react with 2.0 M HCl to produce 12.0 mL of H ₂ gas in 30.0 seconds.
Calculate the rate of reaction
a) In mL of H ₂ /second 12.0 mL H ₂ = 0.400 mL H ₂ /s 30.0 s
30.05 Answer 0.400 mL Hg/s
b) In moles of Hy/second
0.400 m L Hz × 1 L : 1 mol = 1.79 × 10 5 mol Hz/s Answer 1.79 × 10 5 mol Hz/s
1.79×10 5 mol Hz/s

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R Sandford

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- 1. Understanding the eBook Modeling Chemistry U7 Ws 1 V2
 - The Rise of Digital Reading Modeling Chemistry U7 Ws 1 V2
 - o Advantages of eBooks Over Traditional Books
- 2. Identifying Modeling Chemistry U7 Ws 1 V2
 - $\circ \ \ \textbf{Exploring Different Genres}$
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry U7 Ws 1 V2
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Modeling Chemistry U7 Ws 1 V2

- Personalized Recommendations
- Modeling Chemistry U7 Ws 1 V2 User Reviews and Ratings
- Modeling Chemistry U7 Ws 1 V2 and Bestseller Lists
- 5. Accessing Modeling Chemistry U7 Ws 1 V2 Free and Paid eBooks
 - Modeling Chemistry U7 Ws 1 V2 Public Domain eBooks
 - Modeling Chemistry U7 Ws 1 V2 eBook Subscription Services
 - Modeling Chemistry U7 Ws 1 V2 Budget-Friendly Options
- 6. Navigating Modeling Chemistry U7 Ws 1 V2 eBook Formats
 - o ePub, PDF, MOBI, and More
 - Modeling Chemistry U7 Ws 1 V2 Compatibility with Devices
 - Modeling Chemistry U7 Ws 1 V2 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry U7 Ws 1 V2
 - Highlighting and Note-Taking Modeling Chemistry U7 Ws 1 V2
 - Interactive Elements Modeling Chemistry U7 Ws 1 V2
- 8. Staying Engaged with Modeling Chemistry U7 Ws 1 V2
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry U7 Ws 1 V2
- 9. Balancing eBooks and Physical Books Modeling Chemistry U7 Ws 1 V2
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry U7 Ws 1 V2
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry U7 Ws 1 V2
 - Setting Reading Goals Modeling Chemistry U7 Ws 1 V2
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry U7 Ws 1 V2

- Fact-Checking eBook Content of Modeling Chemistry U7 Ws 1 V2
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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