

Answer all the questions in the spaces provided.

- 1 (a) The solubility of the Group 2 sulfates decreases down the group.

Explain this trend.

.....

.....

.....

.....

..... [3]

- (b) Describe what is observed when magnesium and barium are reacted separately with an excess of dilute sulfuric acid.

magnesium

barium

[1]

- (c) The solubility product, K_{sp} , of BaSO_4 is $1.08 \times 10^{-10} \text{ mol}^2 \text{ dm}^{-6}$ at 298 K.

Calculate the solubility of BaSO_4 in g per 100 cm^3 of solution.

solubility of BaSO_4 = g per 100 cm^3 of solution [2]

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Table of Contents Modeling Chemistry U6 Test

1. Understanding the eBook Modeling Chemistry U6 Test
 - The Rise of Digital Reading Modeling Chemistry U6 Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry U6 Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry U6 Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry U6 Test
 - Personalized Recommendations

- Modeling Chemistry U6 Test User Reviews and Ratings
- Modeling Chemistry U6 Test and Bestseller Lists
- 5. Accessing Modeling Chemistry U6 Test Free and Paid eBooks
 - Modeling Chemistry U6 Test Public Domain eBooks
 - Modeling Chemistry U6 Test eBook Subscription Services
 - Modeling Chemistry U6 Test Budget-Friendly Options
- 6. Navigating Modeling Chemistry U6 Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Modeling Chemistry U6 Test Compatibility with Devices
 - Modeling Chemistry U6 Test Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry U6 Test
 - Highlighting and Note-Taking Modeling Chemistry U6 Test
 - Interactive Elements Modeling Chemistry U6 Test
- 8. Staying Engaged with Modeling Chemistry U6 Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry U6 Test
- 9. Balancing eBooks and Physical Books Modeling Chemistry U6 Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry U6 Test
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry U6 Test
 - Setting Reading Goals Modeling Chemistry U6 Test
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry U6 Test
 - Fact-Checking eBook Content of Modeling Chemistry U6 Test

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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