#### Level - I (Based on NCERT)

 Which is not an example of exothermic reaction: (a) N<sub>2</sub> + O<sub>2</sub> → 2NO - Heat (b) N<sub>2</sub> + 3H<sub>2</sub> →2NH<sub>3</sub> + Heat (c) H<sub>2</sub> + I<sub>3</sub> →2HI + Heat (d) All of these 2. Which is an example of endothermic reaction: (a) N<sub>2</sub> + O<sub>2</sub> →2NO - Heat (b) CaCO<sub>3</sub> → CaO + CO<sub>2</sub> − Heat (c) 2KClO<sub>3</sub> → 2kCl + 3O<sub>2</sub> − Heat (d) All of these Which is an example of combination reaction: (a) N<sub>2</sub> + Cl<sub>2</sub> → 2HCl (b) C + O₂ → CO₂ (d) All of these (c)  $N_2 + 3 H_2 \rightarrow 2NH_3$ 4. Which is an example of decomposition reaction? (a) 2Pb(NO<sub>3</sub>)<sub>2</sub> → 2PbO + 4NO<sub>2</sub> + O<sub>2</sub> (b) 2FeSO<sub>4</sub> → Fe<sub>2</sub>O<sub>4</sub> + SO<sub>2</sub> + SO<sub>3</sub> (c) 2H2O Electricity 2H2+ O2 (d) All of these 5. Which is an example of thermal decomposition? (b) 2Al<sub>2</sub>O<sub>3</sub> Electricity 4Al + 3O<sub>3</sub> (a) 2NaCl Electricity 2Na + Cla (c) CaCO<sub>3</sub> → CaO + CO<sub>2</sub> (d) NaOH + HCl →NaCl + H<sub>2</sub>O 6. Which is an example of electrolysis decomposition? (a) 2KClO<sub>3</sub> → 2KCl + 3O<sub>2</sub> (b) BaCl<sub>2</sub> + Na<sub>2</sub>SO<sub>4</sub> → BaSO<sub>4</sub> + 2NaCl (d) 2Al<sub>2</sub>O<sub>2</sub>(I) Electricity 4Al + 3O. (c) 2AgCl hv2Ag + Cl2 7. Which is an example of photo chemical reactions? (a) H<sub>2</sub> + C12 Ker 2 HC1 (b) 2AgBr W 2Ag + Br2 (c) 6CO2 + 6H2O W C6H12O6 (d) All of these 8. Which compound is used in photography? (b) AgCI (c) Na<sub>2</sub>CO<sub>2</sub> (d) (a) and (b) 9. Which is an example of double displacement reaction? (a) BaCl<sub>2</sub> + Na<sub>2</sub>SO<sub>4</sub> → BaSO<sub>4</sub> + 2NaCl (b) AgNO<sub>3</sub> + NaCl →AgCl + NaNO<sub>3</sub> (c) NaOH + HCl →NaCl + H<sub>2</sub>O (d) All of these 10. Fe<sub>2</sub>O<sub>3</sub> + 2Al →Al<sub>2</sub>O<sub>3</sub> + 2Fe is an example of (a) Combination reaction (b) Decomposition reaction (c) Double displacement reaction (d) Displacement reaction 11. Which is an example of double displacement reaction? (a) Mg + 2HCl → MgCl<sub>2</sub> + H<sub>2</sub> (b) ZnCO<sub>3</sub> → ZnO + CO<sub>2</sub> (c) H<sub>2</sub> + Cl<sub>2</sub> →2HCl (d) BaCl<sub>2</sub> + K<sub>2</sub>SO<sub>4</sub> → BaSO<sub>4</sub>2KCl 12. H<sub>2</sub>S + Cl<sub>2</sub> →2HCl + S is this reaction 'Cl<sub>2</sub>' act as: (a) Oxidising agent (b) Reducing agent (d) None of these (c) Both (a) & (b) Which statement is correct regarding the reaction 2PbO + C → C<sub>2</sub>Pb + CO<sub>2</sub> (a) Lead is getting reduced (b) Carbon dioxide is getting oxidised (c) Lead oxide is getting reduced (d) All of these 14. In the reaction MnO₂ + 4HCl → MnCl₂ + 2H2O + Cl₂, MnO₂ act as? (a) Reducing agent (b) Oxidising agent (c) Both (a) & (b) (d) None of these 15. Which is an example of redox reaction? (a) Zn + CuSO<sub>4</sub> → ZnSO<sub>4</sub> + Cu (b) H<sub>2</sub>S + I2 → 2HCl + S (c) CuO + H₂ → Cu + H₂O (d) All of these 16. Which is a formula of rust? (a) Fe<sub>2</sub>O<sub>3</sub> (b) Fe<sub>2</sub>O (c) Fe<sub>2</sub>O<sub>3</sub>×H<sub>2</sub>O (d) FeCO<sub>1</sub> 17. Rusting of Iron is an example: (a) Oxidation (b) Reduction (c) Redox (d) All of these 18. Antioxidant is/are: (d) None of these (a) Oxidising agent (b) Reducing agent (c) Both (a) & (b) 19. Rancidity can be prevented by: (a) Adding antioxidants (b) Storing food in air light

# **Modeling Chemestry Unit 10 Ws 2 Answers**

**Michael Seilmaier** 

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