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MARKING GUIDELINE

NATIONAL CERTIFICATE

JUNE EXAMINATION

**ENTREPRENEURSHIP AND BUSINESS MANAGEMENT N4
(First Paper)**

1 JUNE 2016

This marking guideline consists of 3 pages.

Memorandum For Entrepreneurship And Business

AN Whitehead



Memorandum For Entrepreneurship And Business:

Whispering the Secrets of Language: An Mental Journey through **Memorandum For Entrepreneurship And Business**

In a digitally-driven world wherever monitors reign great and immediate interaction drowns out the subtleties of language, the profound strategies and mental nuances concealed within phrases frequently move unheard. Yet, set within the pages of **Memorandum For Entrepreneurship And Business** a charming literary prize pulsing with raw emotions, lies an extraordinary quest waiting to be undertaken. Composed by a skilled wordsmith, that charming opus invites readers on an introspective trip, softly unraveling the veiled truths and profound influence resonating within the very material of each word. Within the mental depths of the touching review, we shall embark upon a heartfelt exploration of the book is primary subjects, dissect its fascinating writing model, and fail to the strong resonance it evokes strong within the recesses of readers hearts.

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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \quad \Delta H = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta H = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g}) \quad \Delta H = 283 \text{ kJ}$
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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l}) \quad \Delta H = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l}) \quad \Delta H = 1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{ O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta H = 285 \text{ kJ/mole}$

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