DAT General Chemistry Equation Sheet

General Concepts

Density: (d = m/v) Dilutions: $M_1V_1 = M_2V_2$ (or $C_1V_1 = C_2V_2$) % Error Calculations % Error= Actual-theoretical theoretical. Spectrophotometer: Abs = EclAbs = Absorbance c= molar extinction coefficient (molar absorptivity) I = path length

Atomic Structure

Absorption/Emission Line Spectra: $\Delta E = E_{photon}$ Energy of a photon: $E_{photon} = hf = \frac{hc}{r}$ C=3.0+10* PR #/ A-6.63 - 10-34 f-s Photoelectric Effect: $KE_{electron} = E_{photon} - \phi$ downork function (minimum energy to ionize electron)

Chemical Kinetics and Equilibrium

Osmotic Pressure: x=iMRT

General Rate Law: $rate=k[A]^m[B]^n$ Arrhenius Equation: k=Ae-Ea NF $E_{\infty} = activation$ emerger. Equilibrium Constant: 86 - 0.08273 L artmy/mod/ K $K_c = \frac{[products]}{[reactants]}$

Concentration Molarity (M) = moles popular	and the second second second
Lautution	$K_F/K_F = F.P/B.P$ degression/elevation
modes	cometant
molality (m) = kg setem	X _A = mole fraction of A i = wan't Hoff Factor
	# = 0.0821 Latm/mol K
Henry's Law: $P_A = k_H[A]$	M = molarity, m = molality
Colligative Properties	
Freezing Point Depression: &	$T_F = -iK_F m$
Boiling Point Elevation: ΔT_B	$=iK_{B}m$
Raoult's Law (* = pure): P_A =	Y.P.*

Gases/Gas Laws

Awarage Kinetic Energyes

Root-mean-square speed:	$p = \sqrt{\frac{M_m}{M_m}}$
Gas Laws	
Ideal Gas Law: $PV=nRT$ Boyle's Law: $P \propto 1$	-
Charles' Law: V ∝ T	
Avogadro's Law: V ≪ n	
Combined Gas Law: P ₁ V ₁	$\frac{1}{n_1 T_1} = \frac{P_2 V_2}{n_2 T_2}$
Dalton's Law of Partial Pres	isures:
$P_{Tot} = P_A + P_B + \cdots$ $P_A = \chi_A \cdot P_{Tot}$ $\chi_A = mol\ fraction\ A$	R = 0.08211 Latery/mod K n = moles M _m = moler mass
Grahams Law of Effusion:	r = rate of effusion a and b are constants specific to a given gas

Acid-Base

pH Scale $pH = -\log \{H^*\} \{H^*\} = 10^{-pW}$	
$pOH = -\log [OH^-][OH^-] = 10^{-pH}$	
$pH+pOH=14[H^+][OH^-]=1*10^{-14}$	
Weak Acids: $K_a = \frac{[H_2O^+][A^-]}{[HA]}[H^+] = \sqrt{K_a[HA]}$	
Weak Bases: $K_b = \frac{[OH^-][HA]}{[A^-]} [OH^-] = \sqrt{K_b}[B]$	
$K_{ai} \cdot K_{bi} = K_{ac}$	
Buffer: $pH=pK_a + \log \frac{[A^*]}{[BA]}$	
Titration/Neutralization: $n_A M_A V_A = n_B M_B V_B$	

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