

Unit 8 Worksheet 1: Mole relationships

For each of the problems below:

- Write the balanced chemical equation.
- Identify what is given (with units) and what you want to find (with units).
- Use coefficients from balanced equation to determine mole ratio.
- Show set up (organize it!).

- Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce sulfur dioxide and water. How many moles of oxygen gas would be needed to completely burn 8 moles of hydrogen sulfide?

Equation:	<u>2</u> $\text{H}_2\text{S}_{(g)} + \text{3} \text{O}_{2(g)} \rightarrow \text{2} \text{SO}_{2(g)} + \text{2} \text{H}_2\text{O}_{(g)}$				
Before	<u>2</u>	<u>ex</u>	<u>0</u>	<u>0</u>	3.0 mol O ₂
Change	<u>-2</u>	<u>-3</u>	<u>2</u>	<u>2</u>	
After	<u>0</u>	<u>ex</u>	<u>2</u>	<u>2</u>	

- Propane, C₃H₈, burns in air to form carbon dioxide and water. If 12 moles of carbon dioxide are formed, how many moles of propane were burned?

Equation: $2\text{Al} + 6\text{HCl} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2$

Before	0.5	<u>ex</u>	<u>0</u>	<u>0</u>	1.50mol HCL
Change	-0.5	-1.5	+0.5	+0.5	
After	<u>0</u>	<u>ex</u>	<u>0.5</u>	<u>0.75</u>	

- Ammonia, NH₃, for fertilizer is made by causing hydrogen and nitrogen to react at high temperature and pressure. How many moles of ammonia can be made from 0.15 moles of nitrogen gas?

Equation: $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$

Before	<u>ex</u>	<u>1.25</u>	<u>0</u>	2.50 mole
Change	-3.75	-1.25	+25	
After	<u>ex</u>	<u>0</u>	<u>2.5</u>	

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Table of Contents Modeling Chemistry Unit 8 Answers

1. Understanding the eBook Modeling Chemistry Unit 8 Answers
 - The Rise of Digital Reading Modeling Chemistry Unit 8 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Unit 8 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry Unit 8 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Unit 8 Answers
 - Personalized Recommendations
 - Modeling Chemistry Unit 8 Answers User Reviews and Ratings
 - Modeling Chemistry Unit 8 Answers and Bestseller Lists
5. Accessing Modeling Chemistry Unit 8 Answers Free and Paid eBooks
 - Modeling Chemistry Unit 8 Answers Public Domain eBooks
 - Modeling Chemistry Unit 8 Answers eBook Subscription Services

- Modeling Chemistry Unit 8 Answers Budget-Friendly Options
- 6. Navigating Modeling Chemistry Unit 8 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Modeling Chemistry Unit 8 Answers Compatibility with Devices
 - Modeling Chemistry Unit 8 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry Unit 8 Answers
 - Highlighting and Note-Taking Modeling Chemistry Unit 8 Answers
 - Interactive Elements Modeling Chemistry Unit 8 Answers
- 8. Staying Engaged with Modeling Chemistry Unit 8 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry Unit 8 Answers
- 9. Balancing eBooks and Physical Books Modeling Chemistry Unit 8 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry Unit 8 Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry Unit 8 Answers
 - Setting Reading Goals Modeling Chemistry Unit 8 Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry Unit 8 Answers
 - Fact-Checking eBook Content of Modeling Chemistry Unit 8 Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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