

Chemistry Unit 3 Test Review

5.1 Mole and Avogadro's Constant

A **particle** is very small and can be:

- A single atom (Ca)
- A molecule with 2 or more atoms (molecular compound – H₂O)
- An ion with a positive or negative charge (CO₃²⁻)
- A formula unit with two or more ions (ionic compound – NaCl)

The Mole

- Chemists use the mole to represent a specific number of particles
- The number of particles it represents is called the Avogadro constant (6.02×10^{23} particles/mol)
- The unit for mole is mol

Definitions

1. Mole: One mole of any entity is 6.02×10^{23} of the particles (atom, molecule, ion formula unit); the symbol is n; the unit is mol
2. Avogadro's constant: The number of particles in one mole, 6.02×10^{23} ; the symbol is N_A; the unit is particles/mol
3. Particles:
 - Atom or element (Eg. Zn)
 - Molecule or covalent compound (Eg. H₂O)
 - Ions or monatomic/polyatomic ions (Eg. Zn⁺² and CO₃²⁻)
 - Formula unit or ionic compound (Eg. LiCl)

5.2 Molar Mass

Basics: The Mass and The Mole

- The mass of carbon-12 is 12 atomic mass unit (12 u). Every other element has mass relative to one atom of carbon-12
- 1 mole of carbon-12 = 6.02×10^{23} atoms of C-12 = 12g
- The average atomic mass of any element, in atomic mass units, is numerically equivalent to the mass of one mole of atoms (6.02×10^{23} atoms) of that element

Molar Mass: The mass of one mole of its particle – atoms, molecules, ions, or formula units, in grams per mole (g/mol)

6.1 Percentage Composition

Law of Definite Proportions:

- Elements in a chemical compound are always present in the same proportions by mass
- *Law of constant composition:* all molecules or formula units of a compound will have the same chemical composition

Modeling Chemistry Unit 3 Test Answers

Ian Pickup



Modeling Chemistry Unit 3 Test Answers:

Modeling Chemistry Unit 3 Test Answers Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Modeling Chemistry Unit 3 Test Answers**," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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