

$$\text{percent yield} = \frac{\text{actual yield (g)}}{\text{theoretical yield (g)}} \times 100\%$$

↳ from B-E-A chart

Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- Write the balanced chemical equation.
- Identify the given (with units) and what you want to find (with units).
- Show set up with units. Give final answer with units and label.

1. Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?



1 mol H_2O : 18.02 g/mol

36 g $\frac{1\text{ mol H}_2\text{O}}{18.02\text{ g}} = 1.99\text{ mol H}_2\text{O}$

1 mol H_2 : 2.02 g

2 mol $\text{H}_2 = \boxed{4.04\text{ g H}_2}$

2. Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?



1 mol $\text{NaCl} = 58.44\text{ g}$
 $581\text{ g NaCl} \times \frac{1\text{ mol NaCl}}{58.44\text{ g}} = 9.94\text{ mol NaCl}$

1 mol $\text{Na} = 22.99\text{ g}$

$9.94\text{ mol Na} \times 22.99\text{ g} = \boxed{228.52\text{ g Na}}$

Modeling Chemistry U8 V2 Answers

Emilie Sanchez

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Modeling Chemistry U8 V2 Answers:

This book delves into Modeling Chemistry U8 V2 Answers. Modeling Chemistry U8 V2 Answers is an essential topic that needs to be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Modeling Chemistry U8 V2 Answers, encompassing both the fundamentals and more intricate discussions.

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 - Chapter 2: Essential Elements of Modeling Chemistry U8 V2 Answers
 - Chapter 3: Modeling Chemistry U8 V2 Answers in Everyday Life
 - Chapter 4: Modeling Chemistry U8 V2 Answers in Specific Contexts
 - Chapter 5: Conclusion
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