

Chemistry – Unit 5 Worksheet 1

Relative Abundance of the Elements

Relative Mass From Counts

We have established that the relative abundance of an element is the abundance of that isotope expressed as a ratio.

1. Equal abundance of gases means the same number of molecules in the same physical space.
2. Abundance of isotopes is related to mass giving us their relative abundance.

$\text{He}, \text{Ne}, \text{Ar}, \text{Kr}, \text{Xe}, \text{Rn}$

Using the periodic table and relative abundance information find the density of oxygen gas at standard temperature and pressure is 1.43 g/l. Oxygen characterises many substances gas molecules are made up of 2 oxygen atoms. Find the mass of one molecule of oxygen then convert to the mass of hydrogen.

Oxygen
2 p.m.s
1 liter



Relative
Density = $\frac{1.43 \text{ g O}}{1 \text{ liter}}$
Hydrogen
2 p.m.s
1 liter

Hydrogen is the lightest element and therefore was able to demonstrate the relative masses of all the elements using this technique. Measurement of the density of the gases placed at many other elements was difficult at first especially fluorine. We are going to use the relative abundance method to find out the relative mass of fluorine.

Relative Mass From Compounds

Many compounds contain both oxygen and hydrogen and so we can use the relative abundance of these two elements to find the relative mass of the compound. The relative mass of the compound is the sum of the relative masses of the elements in the compound.

For example, the relative mass of water is 18.02 g/mol. The relative mass of oxygen is 16.00 g/mol and the relative mass of hydrogen is 1.01 g/mol. The relative mass of water is 18.02 g/mol. The relative mass of water is 18.02 g/mol.



Modeling Chemistry U5 Ws1 V2 Answers

G Orfield



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