

Learning Objective: Predict bond type and use the VSEPR theory to predict molecular shapes.

1. Identify the following as either having a polar, nonpolar or ionic bond:

a. What are the EN- diff ranges for determining bond types?

Nonpolar $\square$ no diff in EN-, Polar $\square$ 0.1-2.0 diff in EN-, Ionic $\square$ 2.1+ diff in EN-

b. FeO

Polar
(chart)

b. KI

ionic

c. I_2

nonpolar

d. CS

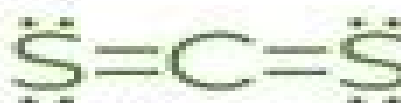
nonpolar (see polar dep. on

2. Draw the Lewis structure of:

a. NH_3



b. CS_2



c. H_2Te
(6e-)



d. BBr_3 (B is stable with

3. Draw the Lewis Structure AND predict the molecular shape of:

a. NO_2^+



Linear

b. CN^+



Bent

c. ClO_3^+



Trigonal Pyramid

d. ClO_2^+



Bent

4. Explain why NO_2^+ and ClO_2^+ have the same molecular geometries even though they have different electron pair geometries. (Lewis structures may be used to supplement your explanation, but in no case will a drawing alone suffice.) (Hint - USE VOCAB!)

NO_2^+ & ClO_2^+ have the same molecular structure shape because they both have two bonded regions of electrons on the central atom, in addition to the presence of lone e- on the central atom that push the bonding regions away.

Modeling Lewis Dot Structures Answer Key

CL Gary



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