

Chemistry – Unit 6 Notes 1 - Thomson Model of the Atom

J. J. Thomson performed experiments with cathode rays in an attempt to understand electricity – which was still a mystery in the late 1800s. Use the website *A Look Inside the Atom*¹ to find the conclusions that Thomson and other physicists drew regarding the mysterious cathode rays.

Thomson's 1897 Experiments - state the conclusions Thomson drew from each of his famous cathode ray experiments:

1. **First Experiment:** Thomson directed the beam at an electrometer and tried to separate the evidence of charge from the path of the beam. *What connection did Thomson find between charge and the cathode rays? Was the charge positive or negative?*
2. **Second Experiment:** Thomson tried passing the cathode ray through an electric field. *How did cathode ray beam behave when it passed through an electric field? What did he conclude after his second experiment?*
3. **Third Experiment:** Thomson did some careful measurements on how much the path of the cathode ray was bent in a magnetic field and how much energy they carried. From this work Thomson could describe the mass/charge ratio of the cathode ray particles. *What amazing result did Thomson find?*

Thomson's Atomic Model: Thomson presented three hypotheses from his experiments. Only two were accepted by physicists – in fact the third was shown to be wrong! From the first two came a model of the atom known as the *Plum Pudding* model. Complete the atom drawing below to illustrate Thomson's plum pudding model. Explain how this fits with his observations.



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Tobias Bleicker



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Table of Contents Modeling Chemistry U6 Review Vanswers

1. Understanding the eBook Modeling Chemistry U6 Review Vanswers
 - The Rise of Digital Reading Modeling Chemistry U6 Review Vanswers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry U6 Review Vanswers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry U6 Review Vanswers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry U6 Review Vanswers
 - Personalized Recommendations
 - Modeling Chemistry U6 Review Vanswers User Reviews and Ratings
 - Modeling Chemistry U6 Review Vanswers and Bestseller Lists

5. Accessing Modeling Chemistry U6 Review Vanswers Free and Paid eBooks
 - Modeling Chemistry U6 Review Vanswers Public Domain eBooks
 - Modeling Chemistry U6 Review Vanswers eBook Subscription Services
 - Modeling Chemistry U6 Review Vanswers Budget-Friendly Options
6. Navigating Modeling Chemistry U6 Review Vanswers eBook Formats
 - ePub, PDF, MOBI, and More
 - Modeling Chemistry U6 Review Vanswers Compatibility with Devices
 - Modeling Chemistry U6 Review Vanswers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry U6 Review Vanswers
 - Highlighting and Note-Taking Modeling Chemistry U6 Review Vanswers
 - Interactive Elements Modeling Chemistry U6 Review Vanswers
8. Staying Engaged with Modeling Chemistry U6 Review Vanswers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry U6 Review Vanswers
9. Balancing eBooks and Physical Books Modeling Chemistry U6 Review Vanswers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry U6 Review Vanswers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modeling Chemistry U6 Review Vanswers
 - Setting Reading Goals Modeling Chemistry U6 Review Vanswers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modeling Chemistry U6 Review Vanswers
 - Fact-Checking eBook Content of Modeling Chemistry U6 Review Vanswers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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