

Ions that form soluble compounds	Exceptions	Ions that form soluble compounds	Exceptions
Group 1 ions (Li^+ , Na^+ , K^+ , ...)	-	Carbonate (CO_3^{2-})	When combined with group 1 ions or ammonium (NH_4^+)
Ammonium (NH_4^+)	-	Chromate (CrO_4^{2-})	When combined with group 1 ions, Ca^{2+} , Mg^{2+} , or ammonium (NH_4^+)
Nitrate (NO_3^-)	-	Phosphate (PO_4^{3-})	When combined with group 1 ions or ammonium (NH_4^+)
Acetate ($\text{C}_2\text{H}_3\text{O}_2^-$ or CH_3COO^-)	-	Sulfide (S^{2-})	When combined with group 1 ions or ammonium (NH_4^+)
Hydrogen Carbonate (HCO_3^-)	-	Hydroxide (OH^-)	When combined with group 1 ions, Ca^{2+} , Ba^{2+} , Sr^{2+} or ammonium (NH_4^+)
Chlorate (ClO_3^-)	-	Sulfates (SO_4^{2-})	When combined with Ag^+ , Ca^{2+} , Sr^{2+} , Ba^{2+} , and Pb^{2+}
Perchlorate (ClO_4^-)	-	Halides (Cl^- , Br^- , I^-)	When combined with Ag^+ , Pb^{2+} , and Hg_2^{2+}

Solubility is important because it directly impacts the amount of radioisotope available for the reaction. It may affect distribution in the radiation source resulting in dosimetry issues. For example, in Benega et. al. [34] a phosphorus-32 radioactive source for spinal cancer treatment was developed by mixing the radioisotope with a catalyst solution. This solution is then added to an epoxy resin. Both solutions need to result in a homogenous product. This won't be achievable if the isotope doesn't properly mix with the catalyst solution (water based).

The solubility product constant, K_{sp} , is the equilibrium constant for a solid substance dissolving in an aqueous solution. It represents the level at which a solute will dissolve in solution. Equation 17 shows the correlation.

$$aA_{(s)} \leftrightarrow bB_{(aq)} + cC_{(aq)} \quad (15)$$

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American Chemical Society

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Model 1 Rules Of Solubility In Aqueous Solutions:

Modeling and Simulation of Reactive Flows A.L. De Bortoli, Greice Andreis, Felipe Pereira, 2015-07-10 Modelling and Simulation of Reactive Flows presents information on modeling and how to numerically solve reactive flows The book offers a distinctive approach that combines diffusion flames and geochemical flow problems providing users with a comprehensive resource that bridges the gap for scientists engineers and the industry Specifically the book looks at the basic concepts related to reaction rates chemical kinetics and the development of reduced kinetic mechanisms It considers the most common methods used in practical situations along with equations for reactive flows and various techniques including flamelet ILDM and Redim for jet flames and plumes with solutions for both In addition the book includes techniques to accelerate the convergence of numerical simulation and a discussion on the analysis of uncertainties with numerical results making this a useful reference for anyone who is interested in both combustion in free flow and in porous media Helps readers learn how to apply applications of numerical methods to simulate geochemical kinetics Presents methods on how to transform the transport equations in several coordinate systems Includes discussions of the basic concepts related to reaction rates chemical kinetics and the development of reduced kinetic mechanisms including the most common methods used in practical situations Offers a distinctive approach that combines diffusion flames and geochemical flow problems

Dense Phase Carbon Dioxide Murat O. Balaban, Giovanna Ferrentino, 2012-04-05 Dense phase carbon dioxide DPCD is a non thermal method for food and pharmaceutical processing that can ensure safe products with minimal nutrient loss and better preserved quality attributes Its application is quite different than for example supercritical extraction with CO₂ where the typical solubility of materials in CO₂ is in the order of 1% and therefore requires large volumes of CO₂ In contrast processing with DPCD requires much less CO₂ between 5 to 8% CO₂ by weight and the pressures used are at least one order of magnitude less than those typically used in ultra high pressure UHP processing There is no noticeable temperature increase due to pressurization and typical process temperatures are around 40 C DPCD temporarily reduces the pH of liquid foods and because oxygen is removed from the environment and because the temperature is not high during the short process time typically about five minutes in continuous systems nutrients antioxidant activity and vitamins are much better preserved than with thermal treatments In pharmaceutical applications DPCD facilitates the production of micronized powders of controlled particle size and distribution Although the capital and operating costs are higher than that of thermal treatments they are much lower than other non thermal technology operations This book is the first to bring together the significant amount of research into DPCD and highlight its effectiveness against microorganisms and enzymes as well as its potential in particle engineering It is directed at food and pharmaceutical industry scientists and technologists working with DPCD and other traditional or non thermal technologies that can potentially be used in conjunction with DPCD It will also be of interest to packaging specialists and regulatory agencies

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,2002 **A Dictionary of Applied Chemistry: Sodium-z** Thomas Edward Thorpe,1913 A Dictionary of applied chemistry v. 5, 1913 Sir Thomas Edward Thorpe,1913 **Inorganic Chemistry** Gary Wulfsberg,2000-03-16 Both elementary inorganic reaction chemistry and more advanced inorganic theories are presented in this one textbook while showing the relationships between the two **Journal of the Chemical Society** ,1918 *Foundations of Inorganic Chemistry* Gary Wulfsberg,2017-11-02 Foundations of Inorganic Chemistry by Gary Wulfsberg is our newest entry into the field of Inorganic Chemistry textbooks designed uniquely for a one semester stand alone course or to be used in a full year inorganic sequence Foundations of Inorganic Chemistry by Gary Wulfsberg is our newest entry into the field of Inorganic Chemistry textbooks designed uniquely for a one semester stand alone course or to be used in a full year inorganic sequence By covering virtually every topic in the test from the 2016 ACS Exams Institute this book will prepare your students for success The new book combines careful pedagogy clear writing beautifully rendered two color art and solved examples with a broad array of original chapter ending exercises It assumes a background in General Chemistry but reviews key concepts and also assumes enrollment in a Foundations of Organic Chemistry course Symmetry and molecular orbital theory are introduced after the student has developed an understanding of fundamental trends in chemical properties and reactions across the periodic table which allows MO theory to be more broadly applied in subsequent chapters Use of this text is expected to increase student enrollment and build students appreciation of the central role of inorganic chemistry in any allied field Key Features Over 900 end of chapter exercises half answered in the back of the book Over 180 worked examples Optional experiments demos Clearly cited connections to other areas in chemistry and chemical sciences Chapter opening biographical vignettes of noted scientists in Inorganic Chemistry Optional General Chemistry review sections Originally rendered two color illustrations throughout *Journal of the American Chemical Society* American Chemical Society,1925 Proceedings of the Society are included in v 1 59 1879 1937 **First Principles of Chemistry** Raymond Bedell Brownlee,1907

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