

Gold Penny Lab

Objective: Students will be observing Metallic Bonding and a single replacement reaction.

Materials:

- Zinc Powder – Zinc can be rinsed with water and used again for this lab.
- 2 M NaOH (Extremely Caustic, Goggles Required)
- Beaker with Water
- Evaporating Dish
- Bunsen Burner



Procedures:

1. Place a few grams of Zinc into an evaporating dish.
2. Add about 10mL of NaOH
3. Slowly heat the solution, be careful not to boil the solution.
4. Add a penny and simmer until the penny becomes silvery. It may be helpful to cover the penny in Zinc. (You may need to flip the penny to cover the sides equally.)
5. Use tongs to pick up the penny and drop it into a beaker of water to rinse it.
6. Hold one edge of the penny with tongs and flip it back and forth in the fire for a few seconds. Repeat until the penny turns gold.
7. Immediately drop it into the water again.

Results:

1. Predict the products and balance the equation: $\text{Cu} + 2\text{NaOH} \rightarrow \text{Zn} + \text{Cu(OH)}_2$
2. What type of reaction occurred when you replaced the Copper with the Zinc? **Single Replacement**
3. What type of bonding was this due to? Defend your answer. **Could defend metallic bonding (due to alloy that is formed) or ionic bonding (due to zinc hydroxide/copper hydroxide ionic).**
4. Name the alloy given to the two metals found on the penny. (You may need to look this up!) **Brass**
5. Pennies made prior to 1982 are made from an alloy composed of 98% Copper and 2% Zinc. Why might pennies be made from an alloy rather than pure copper? **Cost, strength, prevents rust, etc.**
6. If you examined your penny microscopically, do you think the penny would be homogeneous or heterogeneous throughout? **Homogeneous mixture due to the brass alloy**

Gold Penny Lab Answers

RC Schank

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