

Name _____ Class _____ Date _____

Assessment

Chapter Test B

Atomic Physics

MULTIPLE CHOICE

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. What is the frequency of a photon with an energy of $1.00 \times 10^{-19} \text{ J}$?
 $(h = 6.63 \times 10^{-34} \text{ J}\cdot\text{s})$
 - a. $1.00 \times 10^{14} \text{ Hz}$
 - b. $2.00 \times 10^{14} \text{ Hz}$
 - c. $3.00 \times 10^{14} \text{ Hz}$
 - d. $4.00 \times 10^{14} \text{ Hz}$

- _____ 2. Light with an energy equal to three times the work function of a given metal causes the metal to eject photoelectrons. What is the ratio of the maximum photoelectron kinetic energy to the work function?
 - a. 1 : 1
 - b. 2 : 1
 - c. 3 : 1
 - d. 4 : 1

- _____ 3. A monochromatic light beam with a quantum energy value of 3.0 eV is incident upon a photocell. The work function of the photocell is 1.6 eV. What is the maximum kinetic energy of the ejected electrons?
 - a. 4.6 eV
 - b. 4.8 eV
 - c. 1.4 eV
 - d. 2.4 eV

- _____ 4. What causes the bright lines in the emission spectrum of an element to occur?
 - a. Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
 - b. Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
 - c. Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
 - d. Photons are emitted when electrons jump from a lower-energy to a higher-energy state.

- _____ 5. What causes the dark lines in the absorption spectrum of an element to occur?
 - a. Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
 - b. Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
 - c. Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
 - d. Photons are emitted when electrons jump from a lower-energy to a higher-energy state.

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Robert A. Pelcovits, Joshua Farkas



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Table of Contents Motion And Momentum Chapter Test

1. Understanding the eBook Motion And Momentum Chapter Test
 - The Rise of Digital Reading Motion And Momentum Chapter Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Motion And Momentum Chapter Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Motion And Momentum Chapter Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Motion And Momentum Chapter Test
 - Personalized Recommendations
 - Motion And Momentum Chapter Test User Reviews and Ratings

- Motion And Momentum Chapter Test and Bestseller Lists
- 5. Accessing Motion And Momentum Chapter Test Free and Paid eBooks
 - Motion And Momentum Chapter Test Public Domain eBooks
 - Motion And Momentum Chapter Test eBook Subscription Services
 - Motion And Momentum Chapter Test Budget-Friendly Options
- 6. Navigating Motion And Momentum Chapter Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Motion And Momentum Chapter Test Compatibility with Devices
 - Motion And Momentum Chapter Test Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Motion And Momentum Chapter Test
 - Highlighting and Note-Taking Motion And Momentum Chapter Test
 - Interactive Elements Motion And Momentum Chapter Test
- 8. Staying Engaged with Motion And Momentum Chapter Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Motion And Momentum Chapter Test
- 9. Balancing eBooks and Physical Books Motion And Momentum Chapter Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Motion And Momentum Chapter Test
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Motion And Momentum Chapter Test
 - Setting Reading Goals Motion And Momentum Chapter Test
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Motion And Momentum Chapter Test
 - Fact-Checking eBook Content of Motion And Momentum Chapter Test
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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