

CHM 2210 Lab: Melting Point

Introduction

The melting point of a solid is the temperature at which the solid melts (changes to a liquid) at 1 atm of pressure. Melting point is determined by slow heating of a small solid sample placed in a capillary tube. The melting point should be reported as a melting point range: from the temperature at which the first droplet of liquid is observed to the temperature at which the sample becomes a clear liquid, for example: 96.7 – 98.9°C.

A pure solid (pure enough for most purposes) usually has a sharp melting point which means that it melts within a range of 2°C or less.

An impure solid melts at a broader range (3°C or more) and at depressed temperature.

Using melting point to find the identity of a solid is not sufficient because there are too many substances with the same melting points.

However, you can use a mixed melting point as a tool for identification of a solid.

If you mix two different solids with the same melting point, the melting point of the mixture will be broad and depressed. The melting point of a solid and a mixture of this solid with the authentic solid should be the same.

Example:

Measured/Observed melting point of an unknown solid X: 128.0 – 129.5°C

Melting point of known solid A: 130°C

Melting point of known solid B: 158°C

Measured melting point of mixture X + A: 127.5 – 129.0°C

Measured melting point of mixture X + B: 133.5 – 142.0°C

Identity of unknown X: solid A

In the future you will measure melting points of solids mostly to determine their purity.

- **Preparation of Sample**

A solid sample should be dried and crushed with a metal spatula or bottom of a test tube.

If a solid is used for mixed melting point, the 1:1 mixture should be crushed together into a homogeneous mixture.

The pulverized solid should be mounded up and a small (1 – 2 mm) plug of sample should be pushed into the open end of the capillary tube. The sample should be knocked down to the sealed end of the capillary tube by tapping its sealed end on the surface or dropping it into a long glass tubing.

The sample should be 1-2 mm long and packed firmly.

If the sample is too large, the measured melting point may be too high.

If the sample is loosely packed, it may be difficult to determine the melting point range because of the air present.

Melting Point Lab For High School

Joacim Rocklöv

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Gourmet Lab Sarah Reeves Young, 2011 Hands on inquiry based and relevant to every student's life Gourmet Lab serves up a full menu of activities for science teachers of grades 6-12 This collection of 15 hands on experiments each of which includes a full set of both student and teacher pages challenges students to take on the role of scientist and chef as they boil bake and toast their way to better understanding of science concepts from chemistry biology and physics By cooking edible items such as pancakes and butterscotch students have the opportunity to learn about physical changes in states of matter acids and bases biochemistry and molecular structure The Teacher pages include Standards addressed in each lab a vocabulary list safety protocols materials required procedures data analysis student questions answer key and conclusions and connections to spur wrap up class discussions Cross curricular notes are also included to highlight the lesson's connection to subjects such as math and literacy Finally optional extensions for both middle school and high school levels detail how to explore each concept further What better topic than food to engage students to explore science in the natural world

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Chemistry Dorin, Prentice Hall Dictionary Editors, 1992

Michigan School Moderator, 1908

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multiple technologies real time analytics machine learning commodity sensors and embedded systems can serve as tools for learning support and this book details exactly how these powerful tools can be utilized to best effect

Conquering the Electron Derek Cheung, Eric Brach, 2020-02-24 Conquering the Electron offers readers a true and engaging history of the world of electronics beginning with the discoveries of static electricity and magnetism and ending with the creation of the smartphone and the iPad This book shows the interconnection of each advance to the next on the long journey to our modern day technologies Exploring the combination of genius infighting and luck that powered the creation of today s electronic age Conquering the Electron debunks the hero worship so often plaguing the stories of great advances Want to know how AT T s Bell Labs developed semiconductor technology and how its leading scientists almost came to blows in the process Want to understand how radio and television work and why RCA drove their inventors to financial ruin and early graves Conquering the Electron offers these stories and more presenting each revolutionary technological advance right alongside blow by blow personal battles that all too often took place

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exploit it in economic development This updated edition introduces technology management covers the importance of managing information technologies and compares them to existing physical technologies

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