

2

MATTER AND CHANGE

SECTION 2.1 PROPERTIES OF MATTER (pages 39–42)

This section helps you distinguish extensive from intensive properties and identify substances by their properties. It teaches you how to differentiate the three states of matter. It also defines a physical property and lists examples of physical properties and physical changes.

► Describing Matter (page 39)

1. The _____ mass _____ of an object is a measure of the amount of matter the object contains.

2. How does an extensive property differ from an intensive property?

An extensive property depends on the amount of matter, while an intensive property depends on the type of matter.

► Identifying Substances (page 40)

3. Matter that has a uniform and definite composition is called a _____ substance _____.

4. Is the following sentence true or false? All samples of a substance have different physical properties. _____ false _____

5. A physical property is a quality or condition of a substance that can be _____ observed _____ or _____ measured _____ without changing the substance's composition.

6. Circle the letter of the term that is NOT a physical property.

- a. hardness e. boiling point
b. color **d. melting**

7. Look at Table 2.1 on page 40. What is the melting point of bromine? _____ -7°C _____

8. Look at Table 2.1 on page 40. Circle the letter of the substance that is a yellow solid and melts at 115°C.

- a. sulfur**
b. chlorine
c. gold
d. copper

Manual Chemistry Matter And Change Answer Key

**Peter Atkins,Julio de Paula,Ron
Friedman**



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Table of Contents Manual Chemistry Matter And Change Answer Key

1. Understanding the eBook Manual Chemistry Matter And Change Answer Key
 - The Rise of Digital Reading Manual Chemistry Matter And Change Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Manual Chemistry Matter And Change Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Manual Chemistry Matter And Change Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Manual Chemistry Matter And Change Answer Key
 - Personalized Recommendations
 - Manual Chemistry Matter And Change Answer Key User Reviews and Ratings
 - Manual Chemistry Matter And Change Answer Key and Bestseller Lists

5. Accessing Manual Chemistry Matter And Change Answer Key Free and Paid eBooks
 - Manual Chemistry Matter And Change Answer Key Public Domain eBooks
 - Manual Chemistry Matter And Change Answer Key eBook Subscription Services
 - Manual Chemistry Matter And Change Answer Key Budget-Friendly Options
6. Navigating Manual Chemistry Matter And Change Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Manual Chemistry Matter And Change Answer Key Compatibility with Devices
 - Manual Chemistry Matter And Change Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Manual Chemistry Matter And Change Answer Key
 - Highlighting and Note-Taking Manual Chemistry Matter And Change Answer Key
 - Interactive Elements Manual Chemistry Matter And Change Answer Key
8. Staying Engaged with Manual Chemistry Matter And Change Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Manual Chemistry Matter And Change Answer Key
9. Balancing eBooks and Physical Books Manual Chemistry Matter And Change Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Manual Chemistry Matter And Change Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Manual Chemistry Matter And Change Answer Key
 - Setting Reading Goals Manual Chemistry Matter And Change Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Manual Chemistry Matter And Change Answer Key
 - Fact-Checking eBook Content of Manual Chemistry Matter And Change Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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