

Chapter 11: Liquids and Solids

A. Intermolecular Forces

1. The addition of heat changes the state of a compound from solid to liquid to gas as kinetic and potential energy are added to the molecules
 - a. Kinetic Energy involves the motion of the molecules
 - b. Potential Energy involves how far apart the molecules are



2. Intermolecular Forces hold the molecules together
 - a. Ion-Dipole Interactions: Na^+Cl^- interactions with water
 - b. Dipole-Dipole Attractions: interaction of polar molecules
 - c. Hydrogen Bonding: special dipole-dipole attraction in molecules where H is bonded to an electronegative atom (O, N, F, etc...)
 - i. Small size of H lets the dipoles get very close
 - ii. Large electronegativity difference makes attraction great

Intermolecular Forces Projects

H Kauffman



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