

### Worksheet 13 – Intermolecular Forces

Chemical bonds are **intramolecular** forces, which hold atoms together as molecules.

**Intermolecular** forces exist **between** atoms and molecules, determining their **physical** properties; solubility, boiling points, melting points.

Intermolecular forces (IMF) can be ranked qualitatively using Coulomb's Law:

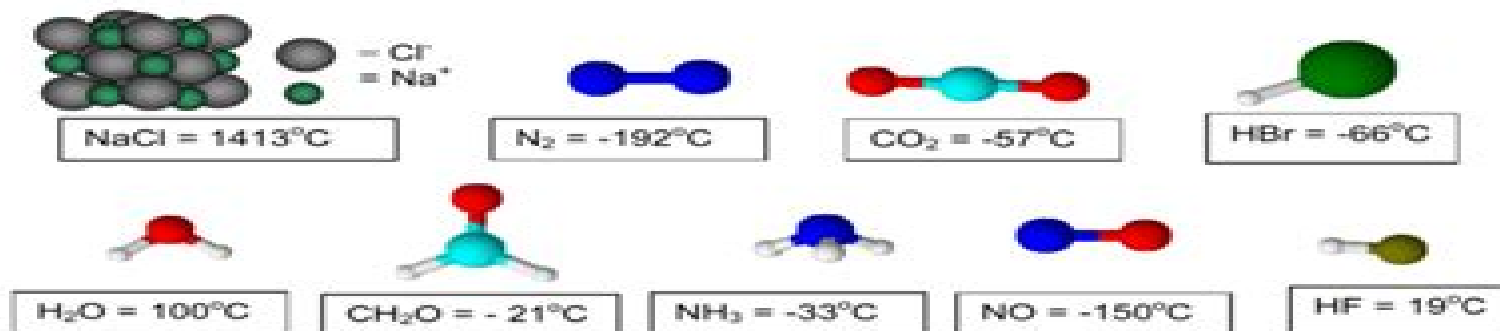
$$\text{force} \propto \frac{Q_1 Q_2}{r}$$

where  $Q_1$  and  $Q_2$  are charges and  $r$  is the distance between the charges.

The consequences are:

- the bigger the charges, the stronger the interaction
- the closer the charges, the stronger the interaction

Shown below are the structures and boiling points of nine chemical species. Use this information to answer the following questions.



1. Which species has the **highest** boiling point? \_\_\_\_\_

This is indicative of the **strongest IMF**.

- How does it differ from all of the other compounds?
- Look at Coulomb's law and identify each variable in the equation to the components of this species.  
 $Q_1$  (positive charge) = \_\_\_\_\_  $Q_2$  (negative charge) = \_\_\_\_\_

This is called an \_\_\_\_\_ interaction (full positive and negative charges).

# Intermolecular Forces Activity

**PT Brinkman**



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