

1-51 The tremendous amount of energy associated with the strong bonds within the nucleus of the atom is called *nuclear energy*. The best-known *fission* reaction involves the split of the uranium atom (the U-235 isotope) into other elements. When a uranium-235 atom absorbs a neutron and splits during a fission process, it produces a cesium-140 atom, a rubidium-93 atom, 3 neutrons, and 3.2×10^{-11} J of energy. Nuclear energy by *fission* is released when two small nuclei combine into a larger one. When two heavy hydrogen (deuterium) nuclei combine during a *fusion* process, they produce a helium-3 atom, a free neutron, and 5.1×10^{-12} J of energy.

1-52 Vast majority of devices, equipment, and appliances used by people operate on electricity. From a thermodynamic point of view, electricity can be converted to other forms of energy (such as mechanical energy and heat) in a 100% conversion efficiency.

1-53 The incorporation of wind power and solar power into the grid involves some irregularities and uncertainties due to changing wind and solar conditions on hourly, daily, and seasonal basis. This requires a more flexible electrical grid system than the existing conventional system in order to accommodate inconsistent supply of renewable electricity. This new grid system is called *smart grid*, which is an important area of research and development for electrical engineers.

1-54 The combustible constituents in the coal are carbon C, hydrogen H_2 , and sulfur S. The higher and lower heating values of hydrogen are 141,800 kJ/kg and 120,000 kJ/kg, respectively and the heating value of carbon is 32,800 kJ/kg (Table A-7). Using their mass fractions, the higher heating value of this particular coal is determined as

$$\begin{aligned} \text{HHV} &= m f_C \times \text{HHV}_C + m f_{H_2} \times \text{HHV}_{H_2} + m f_S \times \text{HHV}_S \\ &= (0.7961)(32,800 \text{ kJ/kg}) + (0.0466)(141,800 \text{ kJ/kg}) + (0.0052)(9160 \text{ kJ/kg}) \\ &= \mathbf{32,770 \text{ kJ/kg}} \end{aligned}$$

Similarly, the lower heating value of the coal is

$$\begin{aligned} \text{LHV} &= m f_C \times \text{LHV}_C + m f_{H_2} \times \text{LHV}_{H_2} + m f_S \times \text{LHV}_S \\ &= (0.7961)(32,800 \text{ kJ/kg}) + (0.0466)(120,000 \text{ kJ/kg}) + (0.0052)(9160 \text{ kJ/kg}) \\ &= \mathbf{31,750 \text{ kJ/kg}} \end{aligned}$$

The difference between the higher and lower heating values is about 3 percent.

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