

Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. The charge on an ion is
a. always positive.
b. always negative.
c. either positive or negative.
d. zero.
- _____ 2. According to the octet rule, a calcium atom has a tendency to
a. lose one electron.
b. lose two electrons.
c. gain one electron.
d. gain two electrons.
- _____ 3. If a compound forms by ionic bonding, which is *not* true?
a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.
b. The net charge of the compound is zero.
c. The compound contains just two atoms, each of opposite charge.
d. Several ions group together in a tightly packed structure.
- _____ 4. The only property listed that is *not* characteristic of ionic compounds is
a. high melting point.
b. hardness.
c. lack of crystal structure.
d. good conductor of electricity.
- _____ 5. Which formula listed below represents a polyatomic ion?
a. NH_4^+
b. H_2SO_4
c. NaCl
d. H_2O
- _____ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because
a. ionic substances tend to vaporize at room temperature.
b. ionic substances are brittle.
c. attractive forces between ions are greater than the attractive forces between molecules.
d. the numbers of positive and negative charges are equal in an ionic compound.

Holt Mcdougal Modern Chemistry Answers

Katrin Zwirgmaier



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Table of Contents Holt Mcdougal Modern Chemistry Answers

1. Understanding the eBook Holt Mcdougal Modern Chemistry Answers
 - The Rise of Digital Reading Holt Mcdougal Modern Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Holt Mcdougal Modern Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Holt Mcdougal Modern Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Holt Mcdougal Modern Chemistry Answers
 - Personalized Recommendations
 - Holt Mcdougal Modern Chemistry Answers User Reviews and Ratings
 - Holt Mcdougal Modern Chemistry Answers and Bestseller Lists
5. Accessing Holt Mcdougal Modern Chemistry Answers Free and Paid eBooks
 - Holt Mcdougal Modern Chemistry Answers Public Domain eBooks
 - Holt Mcdougal Modern Chemistry Answers eBook Subscription Services
 - Holt Mcdougal Modern Chemistry Answers Budget-Friendly Options

6. Navigating Holt Mcdougal Modern Chemistry Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Holt Mcdougal Modern Chemistry Answers Compatibility with Devices
 - Holt Mcdougal Modern Chemistry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Holt Mcdougal Modern Chemistry Answers
 - Highlighting and Note-Taking Holt Mcdougal Modern Chemistry Answers
 - Interactive Elements Holt Mcdougal Modern Chemistry Answers
8. Staying Engaged with Holt Mcdougal Modern Chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Holt Mcdougal Modern Chemistry Answers
9. Balancing eBooks and Physical Books Holt Mcdougal Modern Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Holt Mcdougal Modern Chemistry Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Holt Mcdougal Modern Chemistry Answers
 - Setting Reading Goals Holt Mcdougal Modern Chemistry Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Holt Mcdougal Modern Chemistry Answers
 - Fact-Checking eBook Content of Holt Mcdougal Modern Chemistry Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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