







# Summary of Intermolecular Forces

Force	Model	Basis of Attraction	Energy (kJ/mol)	Example
<b>Nonbonding (Intermolecular)</b>				
Ion-dipole		Ion charge— dipole charge	40–600	$\text{Na}^+ \cdots \text{O} \begin{array}{l} \text{H} \\ \text{H} \end{array}$
H bond		Polar bond to H— dipole charge (high EN of N, O, F)	10–40	$\begin{array}{c} \text{:}\ddot{\text{O}}\text{—H} \\   \\ \text{H} \end{array} \cdots \begin{array}{c} \text{:}\ddot{\text{O}}\text{—H} \\   \\ \text{H} \end{array}$
Dipole-dipole		Dipole charges	5–25	$\text{I—Cl} \cdots \text{I—Cl}$
Ion-induced dipole		Ion charge— polarizable $e^-$ cloud	3–15	$\text{Fe}^{2+} \cdots \text{O}_2$
Dipole-induced dipole		Dipole charge— polarizable $e^-$ cloud	2–10	$\text{H—Cl} \cdots \text{Cl—Cl}$
Dispersion (London)		Polarizable $e^-$ clouds	0.05–40	$\text{F—F} \cdots \text{F—F}$

# Intermolecular Forces Chem 15

**B Lingard**



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