

### Stoichiometry & Percent Yield Practice

- 1) Calculate the molar mass of each compound. Round to 3 decimal places.

a)  $\text{NaCl}$  58.443 g/mol

b)  $\text{FeCl}_3$  162.204 g/mol

c)  $\text{Mg}(\text{NO}_3)_2$  246.055 g/mol

- 2)  $2 \text{AgNO}_3 + \text{Ni} \rightarrow 2 \text{Ag} + \text{Ni}(\text{NO}_3)_2$  BOTH SIGNIFICANTLY

- a) If 6.50 moles of nickel ( $\text{Ni}$ ) reacts with excess silver nitrate ( $\text{AgNO}_3$ ), how many moles of silver ( $\text{Ag}$ ) will be produced?

$$6.50 \text{ mol Ni} \times \frac{2 \text{ mol Ag}}{1 \text{ mol Ni}} = \boxed{13.0 \text{ mol Ag}}$$

mole ratio

- b) If 2.35 moles of silver nitrate ( $\text{AgNO}_3$ ) reacts completely, how many grams of silver ( $\text{Ag}$ ) will be created?

$$2.35 \text{ mol AgNO}_3 \times \frac{2 \text{ mol Ag}}{2 \text{ mol AgNO}_3} \times \frac{107.868 \text{ g Ag}}{1 \text{ mol Ag}} = \boxed{253 \text{ g Ag}}$$

mole ratio

- c) If 2.13 g of  $\text{Ni}$  react with excess silver nitrate, how many moles of  $\text{Ni}(\text{NO}_3)_2$  will be produced?

$$2.13 \text{ g Ni} \times \frac{1 \text{ mol Ni}}{58.693 \text{ g Ni}} \times \frac{1 \text{ mol Ni}(\text{NO}_3)_2}{1 \text{ mol Ni}} = \boxed{0.036 \text{ mol Ni}(\text{NO}_3)_2}$$

mole ratio

- d) If 100 g of  $\text{Ni}$  is produced from the reaction above, what mass of  $\text{Ni}(\text{NO}_3)_2$  will also be produced?

$$100 \text{ g Ni} \times \frac{1 \text{ mol Ni}}{58.693 \text{ g Ni}} \times \frac{1 \text{ mol Ni}(\text{NO}_3)_2}{1 \text{ mol Ni}} \times \frac{246.055 \text{ g Ni}(\text{NO}_3)_2}{1 \text{ mol Ni}(\text{NO}_3)_2} = \boxed{169 \text{ g Ni}(\text{NO}_3)_2}$$

mole ratio

# Holt Stoichiometry Test Answers

**Liying Dong**

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