

Concept

For a gas phase reaction



The equilibrium constant is written as

$$K = \left[\frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b} \right]_{eq}$$

Note: Remember that solids and liquids are not included in the equilibrium constant expression.

Solution



$$K = \left[\frac{(P_{\text{H}_2\text{S}})^8}{(P_{\text{H}_2})^8} \right]_{eq}$$



$$K = \left[\frac{(P_{\text{COCl}_2}) (P_{\text{H}_2})}{(P_{\text{Cl}_2})} \right]$$



$$K = [P_{\text{CO}_2}]_{eq}$$



$$K = \left[\frac{1}{(P_{\text{C}_2\text{H}_2})^3} \right]_{eq}$$

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Rinehart and Winston Staff**



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