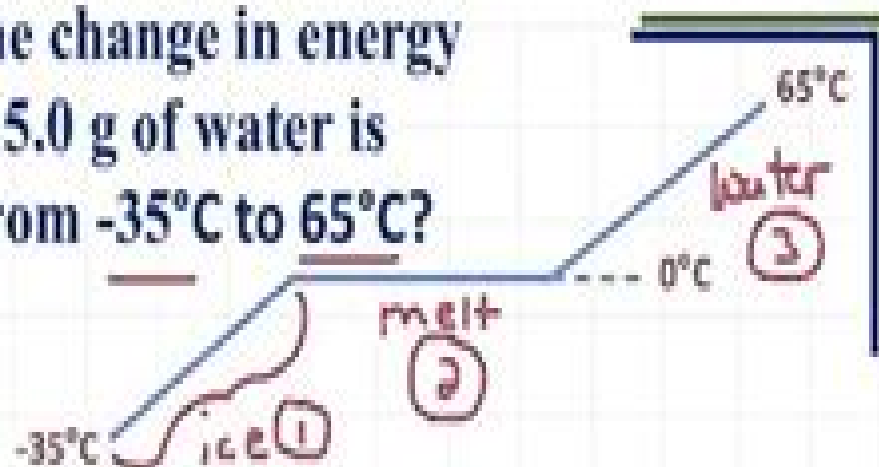


What is the change in energy
when 25.0 g of water is
heated from -35°C to 65°C?



$$\textcircled{1} \quad q = m C_s \Delta T = 25.00 \text{ g} \left(\frac{2.09 \text{ J}}{\text{g}^\circ\text{C}} \right) (0 - (-35))$$

$$q = 1829 \text{ J}$$

$$\textcircled{2} \quad q = n \Delta H_{\text{fus}} = 25.0 \text{ g} \left(\frac{\text{mol}}{18.01 \text{ g}} \right) \left(\frac{6.02 \text{ kJ}}{\text{mol}} \right) (1)$$

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