

General Chemistry 1

Quarter 1, Lesson 14

LIMITING AND EXCESS REAGENTS

Objectives:

- Define the concept of limiting and excess reactants.
- Determine the limiting and excess reactants using mole ratios from balanced chemical equations.

When a chemical reaction is carried out in a flask, the amount of reactants are not always present in the exact proportion stoichiometrically determined from the balanced equation. To make sure that the more expensive reagent is completely used up and converted to the desired product, chemists usually add the cheaper reagent in excess quantity. The reagent that has completely reacted and used up in a reaction is called the **limiting reagent**. The **excess reagent** is the reactant that is present in quantity higher than what is required to react with the limiting reagent.

Sample Problem 1. Let's use the chemical equation below:



Ammonia, NH_3 , is synthesized from the reaction of H_2 and N_2 gases. Suppose 6 moles of H_2 was initially mixed with 4 moles of N_2 gas at high pressure. To determine which of the 2 reactants is the limiting reagent, the amount of NH_3 produced must be computed given the number of moles of H_2 and N_2 and the mole-mole factor from the balanced equation.

$$\text{moles } NH_3 = \cancel{\# \text{ moles of } H_2} \times \frac{2 \text{ moles } NH_3}{\cancel{3 \text{ moles } H_2}}$$

$$\begin{aligned} \text{moles } NH_3 &= 6 \cancel{\text{ moles of } H_2} \times \frac{2 \text{ moles } NH_3}{\cancel{3 \text{ moles } H_2}} \\ \text{moles } NH_3 &= 4 \text{ moles} \end{aligned}$$

$$\text{moles } NH_3 = \cancel{\text{moles of } N_2} \times \frac{2 \text{ moles } NH_3}{\cancel{1 \text{ mole } N_2}}$$

$$\begin{aligned} \text{moles } NH_3 &= 4 \cancel{\text{ moles of } N_2} \times \frac{2 \text{ moles } NH_3}{\cancel{1 \text{ mole } N_2}} \\ \text{moles } NH_3 &= 8 \text{ moles} \end{aligned}$$

From the computed values we determined that if 6 moles of H_2 completely reacts with N_2 , it can produce 4 moles of NH_3 while 4 moles of N_2 can produce 8 moles of NH_3 when fully used up. Since there is less amount of NH_3 produced with 6 moles of H_2 than 4 moles of N_2 , H_2 gas is the limiting reagent while the N_2 gas is the reagent in excess. To determine how much of the 4 moles of N_2 is in excess, we will use the mole-mole factor of N_2 and H_2 .

$$\begin{aligned} \text{moles } N_2 &= 6 \cancel{\text{ moles of } H_2} \times \frac{1 \text{ mole } N_2}{\cancel{3 \text{ moles } H_2}} \\ \text{moles } N_2 &= 2 \text{ moles} \end{aligned}$$

The number of moles of N_2 required to fully react to 6 moles of H_2 is only 2 moles. Thus, the initial 4 moles of N_2 has an excess of 2 moles.

$$\text{excess } N_2 = 2 \text{ moles} - 2 \text{ moles} = 2 \text{ moles}$$

Sample Problem 2. The combustion of ethane produces carbon dioxide and water shown in the reaction below:



- How many moles of CO_2 are produced with 56.2 moles C_2H_6 and 73.4 moles of O_2 ?
Solution:

$$\begin{aligned} \text{moles } CO_2 &= 56.2 \cancel{\text{ moles of } C_2H_6} \times \frac{4 \text{ moles } CO_2}{\cancel{2 \text{ moles } C_2H_6}} \\ \text{moles } CO_2 &= 112.4 \text{ moles} \end{aligned}$$

$$\begin{aligned} \text{moles } CO_2 &= 73.4 \cancel{\text{ moles of } O_2} \times \frac{4 \text{ moles } CO_2}{\cancel{7 \text{ moles } O_2}} \\ \text{moles } CO_2 &= 41.9 \text{ moles} \end{aligned}$$

- Which reagent is the limiting reagent? Which is the excess reagent?
The Limiting reagent is O_2 while the excess reagent is C_2H_6 .
- How many moles of C_2H_6 are in excess?

Limiting And Excess Reactants Answer Key 15

**Nicholas P Cheremisinoff, Avrom
Bendavid-Val**



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web la nourriture du chien n est pas qu une question de choix parmi les innombrables propositions commerciales la bonne alimentation de votre animal nécessite de bien connaître également ses besoins et spécificités vous pourrez ainsi mieux choisir entre les croquettes la pâtée la nourriture maison

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web l alimentation maison prévient l obésité chez le chien du fait que la ration ménagère est riche en eau elle est bénéfique pour le chien car elle favorise la satiété en effet les aliments riches en eau ont tendance à avoir un volume plus important sans apporter un excès de calories

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web jul 3 2017 le chien accepte de manger tous les jours le même aliment et c est meilleur pour sa santé surveiller le poids de son animal c est veiller sur sa santé un chien obèse est un malade en puissance pour lequel il est temps de demander l avis du vétérinaire et de passer à une alimentation allégée avec un volume suffisant pour le rassasier

que mange un chien alimentation du chien planetanimal com - Jun 30 2023

web apr 3 2023 le pourcentage nécessaire de chaque nutriment est différent comme nous le verrons à continuation protéines c est le principal ingrédient de l alimentation d un chien raison pour laquelle vous devriez toujours en céréales le principal céréale inclus les croquettes ou les alimentations

l alimentation du chien conseils pour alimenter son chien - Sep 02 2023

web l alimentation du chiot type d aliments quantité conseils donner à son chien une alimentation de qualité qui lui apporte les vitamines minéraux et autres nutriments essentiels au bon fonctionnement de son organisme est essentiel quel que soit son âge