

Chapter 7, Worksheet 2
AP Chem, Periodic Properties of Elements

Name _____
Period _____

1. As a general rule, the size of atoms increases as you move to the (left/right) or (up/down) on the periodic table. What part of this generalization is counter-intuitive?

2. Predict the atom in each of the following pairs with the larger radius.

- | | | |
|-------------|-------------|-------------|
| a) Be or Mg | b) C or N | c) Se or Cl |
| d) Ba or Y | e) Mg or Sc | f) Te or I |

3. As a general rule, metals that are (farther from/closer to) the boundary between metals and nonmetals have more "metallic character", and nonmetals that are (farther from/closer to) the boundary have more "nonmetallic character".

4. Predict the element with more metallic character.

- | | | | |
|-------------|-------------|------------|-------------|
| a) Cs or Sr | b) Al or Mg | c) Rb or K | d) Rb or Ba |
|-------------|-------------|------------|-------------|

5. Predict the element with more nonmetallic character.

- | | | | |
|-----------|------------|------------|------------|
| a) N or O | b) S or Se | c) As or F | d) B or Si |
|-----------|------------|------------|------------|

6. As a general rule, (metal/nonmetal) ions are larger than (metal/nonmetal) ions.

7. As general rule, metal atoms are (larger/smaller) than their (+) ions, and nonmetal atoms are (larger/smaller) than their (-) ions.

8. In general, the size of either positively charged or negatively charged ions increases as you move to the (left/right) or (up/down) on the periodic table.

9. Predict the larger of each of these pairs of atoms or ions.

- | | | | |
|----------------------------------|---|---|--------------------------------------|
| a) Mg or Mg^{+2} | b) Br or Br^- | c) K or K^+ | d) Na^+ or Mg^{+2} |
| e) Na^+ or F^- | f) Sr^{+2} or Ba^{+2} | g) Mg^{+2} or Al^{+3} | h) I or I^- |
| i) F^- or Cl^- | j) Al^{+3} or P^{-3} | | |

10. As a general rule, first ionization energy (the energy needed to remove the first electron from a neutral gaseous atom) increases as you move to the (left/right) and (up/down) on the periodic table.

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