



Summaries SA

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GRADE 9 NATURAL SCIENCES JUNE TEST

SECTION A

QUESTION 1

1.1 Various options are provided as answers to the following questions. Choose the correct answer and write **ONLY** the LETTER(A-D) next to the QUESTION NUMBER (1.1.1-1.1.10) on your answer sheet, for example 1.1.11 D.

1.1.1 Which of the following can you NOT learn from a chemical formula?

- A Elements in the compound
- B Phase of the element
- C Ratio of atoms
- D Total number of atoms

1.1.2 Which of the following is NOT a sub-atomic particle?

- A Electrons
- B Protons
- C Neutrons
- D Nucleons

1.1.3 What is the prefix for 6 atoms?

- A Hepta-
- B Hexa-
- C Penta-
- D Tri-

1.1.4 An element has an atomic number of 30 and mass number of 65. How many nucleons will it have?

- A 30
- B 35
- C 65
- D 95

1.1.5 Element at group 15, period 2 on the periodic table...

- A Carbon
- B Nitrogen
- C Silicon
- D Oxygen

1.1.6 The following is an example of a semi-metal:

- A Silicon
- B Lithium
- C Magnesium
- D Carbon

1.1.7 Substances on the left side of a chemical equation:

- A Products
- B Reactants
- C Metals
- D Non-metals

1.1.8 Which of the following is NOT a way to prevent rust?

- A Galvanisation
- B Electrolysis
- C Paint
- D Oil or grease

1.1.9 Substance with a pH lower than 7 will be classified as:

- A Lemon Juice
- B Neutral
- C Acidic
- D Alkaline

1.1.10 An neutral substance would turn _____ with universal indicator.

- A Green
- B Violet
- C Colourless
- D Red

[10x1]

1.2 Provide the correct TERM for the following descriptions. Write **ONLY** the correct TERM next to the QUESTION NUMBER (1.2.1-1.2.7) on your answer sheet.

1.2.1 Measure of how strongly acidic or alkaline a substance is.

1.2.2 Reaction of a substance with oxygen.

1.2.3 Name for group 18 on the periodic table.

1.2.4 Process whereby compounds can be broken down using electricity.

1.2.5 H_2 , O_2 , Br_2 , I_2 .

1.2.6 Name for groups 3 to 12 on the periodic table.

1.2.7 When acids and bases react with each other.

[7x1]

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