(40) (10) electrons ... 16 protons 16 ... neutrons ... 16 All correct (1) 11 CHO they are isotopes (1) The word 'isotope' or some derivative is essential (iii) S2-/ sulphide(2-) ion / sulphur -2 ion / sulphur with two extra electrons **CIO**: but not just S or sulphur (Brit.) 610 $X(g) + e \rightarrow X(g) (S(g) + e \rightarrow S(g)) (1)$ $X_{-}(g) + e_{-} \rightarrow X_{-}^{2-}(g) (S_{-}(g) + e_{-} \rightarrow S_{-}^{2-}(g))$ (1) can have Z in place of X2- but not Z2-Penalize omission of state symbol (g) only once. 2 adding a negative onto a negative species requires energy to overcome CHARLE repulsion (1) oursecond electron repelled by negative charge Not second electron repelled by first н 660 (Bdecay from) Y produces only one isotope(39Cl) (1) natural sample contains two isotopes (35 and 37) (1) 2 181 COL magnesium / lead / iron (1) 2. Cab 1 **CHO** iron (1) п carbon / phosphorus (1) 10 CHIEF magnesium / lead (1) CVD carbon / lead (1) Helium (1) because it is the smallest atom (1) with an unshielded nucleus (1) Hydrogen can score max 1 for 'unshielded nucleus' If candidate chooses Carbon, i.e. uses list from part (a), can score Imark for reference to 'smallest atom' 36 deb. ability of a positive ion / cation to attract electrons / distort / pull the electron cloud (1) around an anion (1) size (1) Not atomic radius C 140 F charge (1) charge density worth (2) volume of ion can score in place of size 2 (iii) magnesium chloride magnesium ion smaller than calcium ion (1) polarizes / distorts / attracts electrons around chloride more (1) 2 [14]

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