

Problem Solving *continued*

- 12.** Urea, $(\text{NH}_2)_2\text{CO}$, is an important fertilizer that is manufactured by the following reaction.



What volume of NH_3 at STP will be needed to produce 8.50×10^4 kg of urea if there is an 89.5% yield in the process?

- 13.** An obsolete method of generating oxygen in the laboratory involves the decomposition of barium peroxide by the following equation.



What mass of BaO_2 reacted if 265 mL of O_2 is collected by water displacement at 0.975 atm and $10.^\circ\text{C}$?

- 14.** It is possible to generate chlorine gas by dripping concentrated HCl solution onto solid potassium permanganate according to the following equation.



If excess HCl is dripped onto 15.0 g of KMnO_4 , what volume of Cl_2 will be produced? The Cl_2 is measured at 15°C and 0.959 atm.

- 15.** Ammonia can be oxidized in the presence of a platinum catalyst according to the following equation.



The NO that is produced reacts almost immediately with additional oxygen according to the following equation.



If 35.0 kL of oxygen at STP react in the first reaction, what volume of NH_3 at STP reacts with it? What volume of NO_2 at STP will be formed in the second reaction, assuming there is excess oxygen that was not used up in the first reaction?

- 16.** Oxygen can be generated in the laboratory by heating potassium chlorate. The reaction is represented by the following equation.



What mass of KClO_3 must be used in order to generate 5.00 L of O_2 , measured at STP?

- 17.** One of the reactions in the Solvay process is used to make sodium hydrogen carbonate. It occurs when carbon dioxide and ammonia are passed through concentrated salt brine. The following equation represents the reaction.



- a.** What volume of NH_3 at 25°C and 1.00 atm pressure will be required if 38 000 L of CO_2 , measured under the same conditions, react to form NaHCO_3 ?
- b.** What mass of NaHCO_3 can be formed when the gases in (a) react with NaCl ?

Holt Chemistry Stoichiometry Problem Solving

Jiyuan Zhang



Holt Chemistry Stoichiometry Problem Solving:

The book delves into Holt Chemistry Stoichiometry Problem Solving. Holt Chemistry Stoichiometry Problem Solving is a crucial topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Holt Chemistry Stoichiometry Problem Solving, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Holt Chemistry Stoichiometry Problem Solving
- Chapter 2: Essential Elements of Holt Chemistry Stoichiometry Problem Solving
- Chapter 3: Holt Chemistry Stoichiometry Problem Solving in Everyday Life
- Chapter 4: Holt Chemistry Stoichiometry Problem Solving in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Holt Chemistry Stoichiometry Problem Solving. The first chapter will explore what Holt Chemistry Stoichiometry Problem Solving is, why Holt Chemistry Stoichiometry Problem Solving is vital, and how to effectively learn about Holt Chemistry Stoichiometry Problem Solving.
3. In chapter 2, this book will delve into the foundational concepts of Holt Chemistry Stoichiometry Problem Solving. This chapter will elucidate the essential principles that need to be understood to grasp Holt Chemistry Stoichiometry Problem Solving in its entirety.
4. In chapter 3, this book will examine the practical applications of Holt Chemistry Stoichiometry Problem Solving in daily life. The third chapter will showcase real-world examples of how Holt Chemistry Stoichiometry Problem Solving can be effectively utilized in everyday scenarios.
5. In chapter 4, the author will scrutinize the relevance of Holt Chemistry Stoichiometry Problem Solving in specific contexts. The fourth chapter will explore how Holt Chemistry Stoichiometry Problem Solving is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, this book will draw a conclusion about Holt Chemistry Stoichiometry Problem Solving. The final chapter will summarize the key points that have been discussed throughout the book.
- The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Holt Chemistry Stoichiometry Problem Solving.

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