

Assessment

Chapter Test B

Atomic Physics

MULTIPLE CHOICE

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

1. What is the frequency of a photon with an energy of $1.99 \times 10^{-19} \text{ J}$?
($h = 6.63 \times 10^{-34} \text{ J}\cdot\text{s}$)
a. $1.00 \times 10^{14} \text{ Hz}$
b. $2.00 \times 10^{14} \text{ Hz}$
c. $3.00 \times 10^{14} \text{ Hz}$
d. $4.00 \times 10^{14} \text{ Hz}$
2. Light with an energy equal to three times the work function of a given metal causes the metal to eject photoelectrons. What is the ratio of the maximum photoelectron kinetic energy to the work function?
a. 1 : 1
b. 2 : 1
c. 3 : 1
d. 4 : 1
3. A monochromatic light beam with a quantum energy value of 3.0 eV is incident upon a photocell. The work function of the photocell is 1.6 eV. What is the maximum kinetic energy of the ejected electrons?
a. 4.6 eV
b. 4.8 eV
c. 1.4 eV
d. 2.4 eV
4. What causes the bright lines in the emission spectrum of an element to occur?
a. Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
b. Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
c. Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
d. Photons are emitted when electrons jump from a lower-energy to a higher-energy state.
5. What causes the dark lines in the absorption spectrum of an element to occur?
a. Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
b. Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
c. Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
d. Photons are emitted when electrons jump from a lower-energy to a higher-energy state.

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Lingsheng Yao



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Table of Contents Holt And Technology Introduction To Atoms Chapter Test B

1. Understanding the eBook Holt And Technology Introduction To Atoms Chapter Test B
 - The Rise of Digital Reading Holt And Technology Introduction To Atoms Chapter Test B
 - Advantages of eBooks Over Traditional Books
2. Identifying Holt And Technology Introduction To Atoms Chapter Test B
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Holt And Technology Introduction To Atoms Chapter Test B
 - User-Friendly Interface

4. Exploring eBook Recommendations from Holt And Technology Introduction To Atoms Chapter Test B
 - Personalized Recommendations
 - Holt And Technology Introduction To Atoms Chapter Test B User Reviews and Ratings
 - Holt And Technology Introduction To Atoms Chapter Test B and Bestseller Lists
5. Accessing Holt And Technology Introduction To Atoms Chapter Test B Free and Paid eBooks
 - Holt And Technology Introduction To Atoms Chapter Test B Public Domain eBooks
 - Holt And Technology Introduction To Atoms Chapter Test B eBook Subscription Services
 - Holt And Technology Introduction To Atoms Chapter Test B Budget-Friendly Options
6. Navigating Holt And Technology Introduction To Atoms Chapter Test B eBook Formats
 - ePub, PDF, MOBI, and More
 - Holt And Technology Introduction To Atoms Chapter Test B Compatibility with Devices
 - Holt And Technology Introduction To Atoms Chapter Test B Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Holt And Technology Introduction To Atoms Chapter Test B
 - Highlighting and Note-Taking Holt And Technology Introduction To Atoms Chapter Test B
 - Interactive Elements Holt And Technology Introduction To Atoms Chapter Test B
8. Staying Engaged with Holt And Technology Introduction To Atoms Chapter Test B
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Holt And Technology Introduction To Atoms Chapter Test B
9. Balancing eBooks and Physical Books Holt And Technology Introduction To Atoms Chapter Test B
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Holt And Technology Introduction To Atoms Chapter Test B
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Holt And Technology Introduction To Atoms Chapter Test B
 - Setting Reading Goals Holt And Technology Introduction To Atoms Chapter Test B
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Holt And Technology Introduction To Atoms Chapter Test B
 - Fact-Checking eBook Content of Holt And Technology Introduction To Atoms Chapter Test B
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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