

HEAT AND ITS MEASUREMENT

NAME: _____

Heat for a change in temperature is measured in units of calories or joules. When measuring temperature change, the heat can either be calculated using the formula:

$$Q = mc\Delta T$$

(Q = heat in J or calories, m = mass in grams, c = specific heat capacity, ΔT = change in temperature)

During a phase change, we use the formula:

$$Q = m\Delta H$$

(Q = heat in J or calories, m = mass, ΔH = heat of vaporization or fusion)

Solve the following problems.

1. How many joules of heat are given off when 5.0 g of water cools from 75°C to 25°C? (specific heat of water = 4.18 J/g°C)	-1045 J
2. How many calories are given off by 10 g of water as it cools to 10°C from 25°C? (specific heat of water = 1.00 cal/g°C)	-250 cal
3. How many joules does it take to melt 20 g of ice at 0°C? (heat of fusion = 333 J/g)	6660 J
4. How many calories are given off when 80 g of water cools from 80°C to 20°C? (specific heat of water = 1.00 cal/g°C)	-4800 cal
5. How many joules of heat are necessary to raise the temperature of 25 g of water from 10°C to 25°C?	2625 J
6. How many calories are given off when 50 g of water (50°C) freezes? (heat of fusion = 79.72 cal/g)	-3986 cal

Het And Its Measurement Answer Key

Scott C. Dulebohn



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