The following cell is a chemical cell connected to a light bulb and voltmeter. Salt Bridge Graphite Electrode A., Graphite Electrode B 1.50cmo2 1.50cm3 IM K2Cr2Ov<sub>initi</sub>+ 0.6M Fc(NO<sub>3</sub>)20ea + 0.8M AgNOsimo IM MgSO4ced Draw the direction of electron flow on the figure. (1 mark) (b) For the above setup, Write the half equations for the reaction occur in the two cells. (1 mark) (ii) Write the overall equation for the cells. (I mark) Salt bridge in the cell can be prepared by soaking filter paper into NaCl solution. (c) State the function of salt bridge. (1) (2 marks) (ii) Other than NaCl solution, what else can be used. (1 mark)

# **Galvanic Cell Sample Exam Questions Withe Answers**

**S** Ashworth

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