

2. Magnesium carbonate reacts with hydrochloric acid.



- (a) What mass of magnesium carbonate would react with excess hydrochloric acid to produce  $240\text{cm}^3$  of carbon dioxide, measured at room temperature and pressure?

Data: 1 mol of any gas occupies  $24.0\text{dm}^3$  at room temperature and pressure

Molar mass of magnesium carbonate =  $84.3\text{g mol}^{-1}$

- ☒ A 0.843 g  
☐ B 8.43 g  
☐ C 84.3 g  
☐ D 843 g

$$\begin{aligned}\text{Volume of CO}_2 &= 240\text{cm}^3 = 0.24\text{dm}^3 \\ \text{moles of CO}_2 &= \frac{\text{Volume in dm}^3}{24} = \frac{0.24}{24} = 0.01 \\ \text{MgCO}_3 &\text{ contains same moles as CO}_2 \\ \text{moles} &= \frac{\text{mass}}{\text{Mr}} ; \text{ mass} = \text{moles} \times \text{Mr} \\ &= 0.01 \times 84.3 \\ &= 0.843\text{g}\end{aligned}$$

- (b) What is the minimum mass of magnesium carbonate needed to neutralise  $50.0\text{cm}^3$  of  $0.250\text{mol dm}^{-3}$  hydrochloric acid?

- ☐ A 0.423 g  
☒ B 0.527 g  
☐ C 1.05 g  
☐ D 2.11 g

$$\begin{aligned}\text{Concentration of HCl} &= \frac{\text{moles}}{\text{Vol in dm}^3} \\ \text{moles} &= \text{Conc} \times \text{Vol in dm}^3 \\ &= 0.250 \times 0.05 \\ &= 0.0125\text{ moles of HCl}\end{aligned}$$

$$\begin{aligned}\text{MgCO}_3 : \text{HCl} \\ 1 : 2 \\ \frac{1}{2} : 1 \\ \frac{1}{2} \times 0.0125 : 0.0125 \\ 6.25 \times 10^{-3} : 0.0125 \\ \downarrow \qquad \qquad \downarrow \\ \text{MgCO}_3 \qquad \text{HCl} \\ \text{moles} \qquad \text{moles}\end{aligned}$$

- (c) What would be seen at the end of the reaction with excess acid?

- ☒ A A colourless solution  
☐ B A coloured solution  
☐ C A white precipitate  
☐ D A coloured precipitate

all  $\text{MgCO}_3$  is utilized  
and color of  
 $\text{MgCO}_2$  colorless

$$\text{mass} = \text{moles} \times \text{Mr}$$

(Total for Question 2 = 3 marks)

$$\begin{aligned}&= 6.25 \times 10^{-3} \times 84.3 \\ &= 0.527\text{g}\end{aligned}$$

# Multiple Choice Questions Of Stoichiometry

**S Ashworth**



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## **Table of Contents Multiple Choice Questions Of Stoichiometry**

1. Understanding the eBook Multiple Choice Questions Of Stoichiometry
  - The Rise of Digital Reading Multiple Choice Questions Of Stoichiometry
  - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Questions Of Stoichiometry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Multiple Choice Questions Of Stoichiometry
  - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Questions Of Stoichiometry
  - Personalized Recommendations
  - Multiple Choice Questions Of Stoichiometry User Reviews and Ratings
  - Multiple Choice Questions Of Stoichiometry and Bestseller Lists

5. Accessing Multiple Choice Questions Of Stoichiometry Free and Paid eBooks
  - Multiple Choice Questions Of Stoichiometry Public Domain eBooks
  - Multiple Choice Questions Of Stoichiometry eBook Subscription Services
  - Multiple Choice Questions Of Stoichiometry Budget-Friendly Options
6. Navigating Multiple Choice Questions Of Stoichiometry eBook Formats
  - ePub, PDF, MOBI, and More
  - Multiple Choice Questions Of Stoichiometry Compatibility with Devices
  - Multiple Choice Questions Of Stoichiometry Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Multiple Choice Questions Of Stoichiometry
  - Highlighting and Note-Taking Multiple Choice Questions Of Stoichiometry
  - Interactive Elements Multiple Choice Questions Of Stoichiometry
8. Staying Engaged with Multiple Choice Questions Of Stoichiometry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Multiple Choice Questions Of Stoichiometry
9. Balancing eBooks and Physical Books Multiple Choice Questions Of Stoichiometry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Multiple Choice Questions Of Stoichiometry
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Multiple Choice Questions Of Stoichiometry
  - Setting Reading Goals Multiple Choice Questions Of Stoichiometry
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Multiple Choice Questions Of Stoichiometry
  - Fact-Checking eBook Content of Multiple Choice Questions Of Stoichiometry
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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