

Variable: Residuals

Anderson-Darling Normality Test

A-Squared: 0.463
P-Value: 0.210

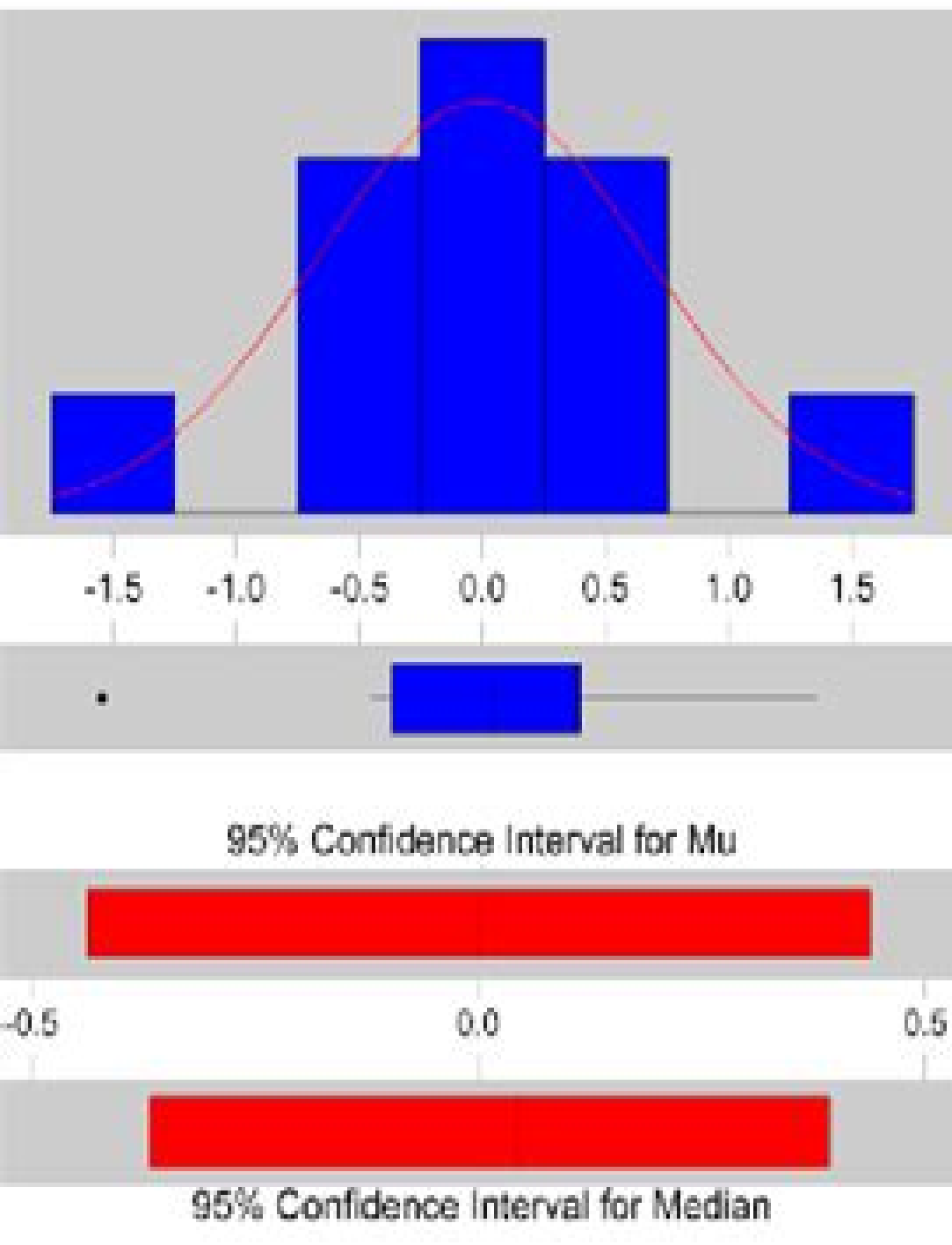
Mean 0.000000
StDev 0.688705
Variance 0.474315
Skewness -3.9E-01
Kurtosis 2.49982
N 12

Minimum -1.54667
1st Quartile -0.36833
Median 0.04333
3rd Quartile 0.39333
Maximum 1.36333

95% Confidence Interval for Mu
-0.43758 0.43758

95% Confidence Interval for Sigma
0.48788 1.16934

95% Confidence Interval for Median
-0.36790 0.39228



Minitab 16 User Guide

R Bogdan



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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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 $\text{hcl aq} + \text{naoh aq} \rightarrow \text{nacl aq} + \text{h}_2\text{O l}$ one mole of hcl would be fully neutralized by one mole of naoh

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide naoh and oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{C}_2\text{O}_4 + 2\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a pH scale is used the pH scale ranges from 0 to 14 a pH of 7 is neutral in an acid base reaction

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction
 $\text{NaOH} + \text{C}_6\text{H}_4\text{COOH} \rightarrow \text{C}_6\text{H}_4\text{COONa} + \text{H}_2\text{O}$ l the point of equivalence is the point of titration when the number of moles of naoh na added is equal to the number of moles of khp nb in the