

82. Additional Questions

1. The percent by mass composition of a salt was found to be 76.58% potassium, 6.48% sulfur, and 16.94% oxygen. What is the empirical formula of this salt?

$$56.58g \text{ of K} \left(\frac{\text{molar K}}{39.098 \text{ g K}} \right) = 1.447 \text{ molar K} \quad \frac{1.447 \text{ molar K}}{0.723 \text{ molar S}} = 2$$

$$6.48g \text{ of S} \left(\frac{\text{molar S}}{32.065 \text{ g S}} \right) = 0.202 \text{ molar S} \quad \frac{0.202 \text{ molar S}}{0.202 \text{ molar S}} = 1$$

$$16.94g \text{ of O} \left(\frac{\text{molar O}}{16.000 \text{ g O}} \right) = 1.059 \text{ molar O} \quad \frac{1.059 \text{ molar O}}{0.202 \text{ molar S}} = 5$$

$$\text{empirical formula} = K_2S_1O_5$$

2. A compound consisting iron and sulfur was formed by combining 3.235 g of iron with 1.426 g of sulfur. What is the empirical formula of the compound?

$$3.235g \text{ of Fe} \left(\frac{\text{molar Fe}}{55.845 \text{ g Fe}} \right) = 0.05794 \text{ molar Fe} \quad \frac{0.05794 \text{ molar Fe}}{0.05794 \text{ molar Fe}} = 1 + 2 = 2$$

$$1.426g \text{ of S} \left(\frac{\text{molar S}}{32.067 \text{ g S}} \right) = 0.04447 \text{ molar S} \quad \frac{0.04447 \text{ molar S}}{0.05794 \text{ molar Fe}} = 0.7675 + 2 = 3$$

$$\text{empirical formula} = Fe_2S_3$$

3. Propylene has a molar mass of 42.08 g/mole and is composed of 84.3% hydrogen and 85.7% carbon. What is the molecular formula of propylene?

$$14.3g \text{ of H} \left(\frac{\text{molar H}}{1.008 \text{ g H}} \right) = 14.15 \text{ molar H} \quad \frac{14.15 \text{ molar H}}{7.134 \text{ molar C}} = 2$$

$$85.7g \text{ of C} \left(\frac{\text{molar C}}{12.011 \text{ g C}} \right) = 7.134 \text{ molar C} \quad \frac{7.134 \text{ molar C}}{7.134 \text{ molar C}} = 1$$

$$\text{empirical formula} = CH_2$$

$$CH_2 = 12.011 + 2(1.008) = 14.027 \quad \frac{42.08}{14.027} = 3$$

$$\text{molecular formula} = C_3H_6$$

Magnesium Oxide Lab Answers

JL Elias

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Table of Contents Magnesium Oxide Lab Answers

1. Understanding the eBook Magnesium Oxide Lab Answers
 - The Rise of Digital Reading Magnesium Oxide Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Magnesium Oxide Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Magnesium Oxide Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Magnesium Oxide Lab Answers
 - Personalized Recommendations
 - Magnesium Oxide Lab Answers User Reviews and Ratings
 - Magnesium Oxide Lab Answers and Bestseller Lists
5. Accessing Magnesium Oxide Lab Answers Free and Paid eBooks
 - Magnesium Oxide Lab Answers Public Domain eBooks

- Magnesium Oxide Lab Answers eBook Subscription Services
- Magnesium Oxide Lab Answers Budget-Friendly Options
- 6. Navigating Magnesium Oxide Lab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Magnesium Oxide Lab Answers Compatibility with Devices
 - Magnesium Oxide Lab Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Magnesium Oxide Lab Answers
 - Highlighting and Note-Taking Magnesium Oxide Lab Answers
 - Interactive Elements Magnesium Oxide Lab Answers
- 8. Staying Engaged with Magnesium Oxide Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Magnesium Oxide Lab Answers
- 9. Balancing eBooks and Physical Books Magnesium Oxide Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Magnesium Oxide Lab Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Magnesium Oxide Lab Answers
 - Setting Reading Goals Magnesium Oxide Lab Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Magnesium Oxide Lab Answers
 - Fact-Checking eBook Content of Magnesium Oxide Lab Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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