KEY

FINAL EXAM PRACTICE SHEET

- Write the molecular, total ionic, and net ionic equations of the following reactions
 (15pts): + predict the products
- a. Zinc metal is placed in aqueous hydrochloric acid

ME
$$Z_{n}(s) + 2H(e_{qq}) \rightarrow Z_{n}(e_{q}) + H_{2}(q)$$

TE $Z_{n}(s) + 2H'e_{q}) + 2CC'e_{q}) \rightarrow Z_{n}^{2+}(e_{q}) + 2CC'(e_{q}) + H_{2}(q)$
NIE $Z_{n}(s) + 2H'e_{q}) \rightarrow Z_{n}^{2+}(e_{q}) + H_{2}(q)$

b. Aqueous barium chloride is mixed with aqueous sodium sulfate

ME Bally + Nay SO4 (an)
$$\rightarrow$$
 BaSO4 (as) + 2 Nall (an)

IE Baten) + 20 (an) + 2 Nat (an) + 50 to (an) \rightarrow BaSO4 (s) + 2 Nat (an)

+ 20 (an)

NIE Baten) + 50 to (an) \rightarrow BaSO4 (s)

c. Aqueous ammonium sulfate is mixed with aqueous potassium hydroxide

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Clifford Lane

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