

# Chemistry Revision

## Moles, Empirical formula, Ar and Mr

### $A_r$ & $M_r$

- Relative atomic mass:
  - the weighted average mass of the naturally occurring isotopes of an element as compared to  $1/12$  of the mass of carbon-12
  - weighted average means it factors in the abundance of each isotope
  - $(\% \times A_r) + (\% \times A_r) / 100$
- Relative formula mass:
  - the sum of all the atomic masses
  - use relative formula mass to work out percentage composition

Question: find percentage by mass of copper in copper(II) oxide,  $\text{CuO}$

$A_r$ : O = 16; Cu = 64

$$\begin{aligned}M_r &= 64 + 16 \\&= 80\end{aligned}$$

$$\begin{aligned}\text{Percentage of copper} &= 64/80 \times 100 \\&= 80\%\end{aligned}$$

### Moles

- The number of carbon atoms in exactly 12g of carbon-12
- Formula:
  - number of moles = mass (g) /  $M_r$
  - number of moles = number of particles /  $6.02 \times 10^{23}$
- Reacting mass calculations:
  - mole (convert mass into moles)
  - mole ratio (from balanced equation)
  - un-mole (convert back to grams)
  - when un-moling in gas calculations, always multiply by 24 instead of  $M_r$

Question: find the mass aluminium oxide when 135g of aluminium is burned with oxygen



- mole = mass/ $M_r$  =  $135/27$  = 5 mol Al
- mole ratio =  $4:2$  =  $2:1$  = 2.5 mol  $\text{Al}_2\text{O}_3$
- un-mole =  $2.5 \times 102$  = 255g  $\text{Al}_2\text{O}_3$

# Igcse Edexcel Chemistry May 2013

**Michal Rosen-Zvi**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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