

KEY

Radioactive Half-Life Problems

1.) Calculate the number of half lives it would take for 100 grams of Radium-223 to decay into 3.125 grams of Radon-219.

$$100 \xrightarrow[11 \text{ days}]{①} 50 \xrightarrow[11 \text{ days}]{②} 25 \xrightarrow[11 \text{ days}]{③} 12.5 \xrightarrow[11 \text{ days}]{④} 6.25 \xrightarrow[11 \text{ days}]{⑤} 3.125$$

55

2.) How long would it take for 100 grams of Radium-223 to decay into 3.125 grams of Radon-219 if the half life of Radium-223 is 11 days?

$$100g \longrightarrow 3.125g \text{ TAKES } 5 \text{ HALF-LIVES}$$

$$5 \times 11 \text{ days} = 55 \text{ DAYS}$$

3.) Astatine-217 has a half life of 20 milliseconds. If you start with a 30 gram sample of Astatine-217 how much will be left after 100 milliseconds? Hint: how many half lives in 100 milliseconds.

$$100 \text{ ms} = 5 \text{ HALF-LIVES}$$

$$20 \overline{) 100} = 5$$

$$30g \xrightarrow{①} 15g \xrightarrow{②} 7.5g \xrightarrow{③} 3.75g \xrightarrow{④} 1.875g \xrightarrow{⑤} .9375g$$

4.) Carbon-14 has a half life of 5730 years. An archeologist discovers a piece of cloth in an old cave. If it could be determined that 3.5 half lives had passed since the cloth was made how old would the cloth be?

$$5730 \text{ YRS} \times 3.5 \text{ H.L.} = 20,055 \text{ years}$$

5.) If you start with a sample of 100 atoms and determine the half life to be 10 days would the half life be any different if you then repeated your measurement using 10,000 atoms? Explain

NO - IT WOULD JUST TAKE LONGER FOR THE SUBSTANCE TO COMPLETELY DECAY.

6.) Thorium-232 decays into Radium-228 with a half-life of 1.9 years. If you start with a 60gram sample of Thorium-232 how much would you have left after 5.7 years?

$$1.9 \overline{) 5.7} = 3 \text{ HALF-LIVES}$$

$$60g \xrightarrow{①} 30g \xrightarrow{②} 15g \xrightarrow{③} 7.5g$$

7.) Lead-212 decays into Bismuth-212 with a half-life of 11 hours. If you start with 200 atoms of Lead-212 how many atoms of Bismuth-212 would you have after 22 hours.

$$11 \overline{) 22} = 2 \text{ HALF-LIVES}$$

$$\text{lead atoms} = 200 \xrightarrow{①} 100 \xrightarrow{②} 50 = \text{lead atoms}, 50 = \text{BISMUTH-212}$$

Half Life Calculation Answer Key Instructional Fair

Scott C. Dulebohn

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Table of Contents Half Life Calculation Answer Key Instructional Fair

1. Understanding the eBook Half Life Calculation Answer Key Instructional Fair
 - The Rise of Digital Reading Half Life Calculation Answer Key Instructional Fair
 - Advantages of eBooks Over Traditional Books
2. Identifying Half Life Calculation Answer Key Instructional Fair
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Half Life Calculation Answer Key Instructional Fair
 - User-Friendly Interface
4. Exploring eBook Recommendations from Half Life Calculation Answer Key Instructional Fair
 - Personalized Recommendations
 - Half Life Calculation Answer Key Instructional Fair User Reviews and Ratings
 - Half Life Calculation Answer Key Instructional Fair and Bestseller Lists
5. Accessing Half Life Calculation Answer Key Instructional Fair Free and Paid eBooks

- Half Life Calculation Answer Key Instructional Fair Public Domain eBooks
 - Half Life Calculation Answer Key Instructional Fair eBook Subscription Services
 - Half Life Calculation Answer Key Instructional Fair Budget-Friendly Options
6. Navigating Half Life Calculation Answer Key Instructional Fair eBook Formats
- ePub, PDF, MOBI, and More
 - Half Life Calculation Answer Key Instructional Fair Compatibility with Devices
 - Half Life Calculation Answer Key Instructional Fair Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Half Life Calculation Answer Key Instructional Fair
 - Highlighting and Note-Taking Half Life Calculation Answer Key Instructional Fair
 - Interactive Elements Half Life Calculation Answer Key Instructional Fair
8. Staying Engaged with Half Life Calculation Answer Key Instructional Fair
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Half Life Calculation Answer Key Instructional Fair
9. Balancing eBooks and Physical Books Half Life Calculation Answer Key Instructional Fair
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Half Life Calculation Answer Key Instructional Fair
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Half Life Calculation Answer Key Instructional Fair
- Setting Reading Goals Half Life Calculation Answer Key Instructional Fair
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Half Life Calculation Answer Key Instructional Fair
- Fact-Checking eBook Content of Half Life Calculation Answer Key Instructional Fair
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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