

# STD XII - CHEMISTRY

Time : 3 hrs.]

MARCH - 2015

[Marks : 70

Note : (i) All questions are compulsory.

(ii) Answer to the two sections should be written in the same answer book.

(iii) Figures to the right hand side indicate full marks.

(iv) Write balanced chemical equations and draw neat and labelled diagrams wherever necessary.

(v) Use of logarithmic table is allowed.

(vi) Answer to every new question must be started on a new page.

## SECTION - I

- Q. 1. Select and write the most appropriate answers from the given alternatives for each sub-question : [7]**
- (i) p-type semi-conductors are made by mixing silicon with impurities of \_\_\_\_\_.  
(a) germanium (b) boron (c) arsenic (d) antimony
- (ii) Amongst the following identify the criterion for a process to be at equilibrium :  
(a)  $\Delta G < 0$  (b)  $\Delta G > 0$  (c)  $\Delta S_{\text{total}} = 0$  (d)  $\Delta S < 0$
- (iii) Colligative property depends only on \_\_\_\_\_ in a solution.  
(a) Number of solute particles. (b) Number of solvent particles.  
(c) Nature of solute particles. (d) Nature of solvent particles.
- (iv) The charge of how many coulomb is required to deposit 1.0 g of sodium metal (molar mass  $23.0 \text{ g mol}^{-1}$ ) from sodium ions is \_\_\_\_\_.  
(a) 2098 (b) 96500 (c) 193000 (d) 4196
- (v) What is the chemical composition of malachite ?  
(a)  $\text{CuO} \cdot \text{CuCO}_3$  (b)  $\text{Cu}(\text{OH})_2 \cdot \text{CuCO}_3$  (c)  $\text{CuO} \cdot \text{Cu}(\text{OH})_2$  (d)  $\text{Cu}_2\text{O} \cdot \text{Cu}(\text{OH})_2$
- (vi) The element that does NOT exhibit allotropy is \_\_\_\_\_.  
(a) As (b) Sb (c) Bi (d) N
- (vii) The integrated rate equation for first order reaction is  $A \xrightarrow{\quad} \text{products}$  \_\_\_\_\_.  
(a)  $K = 2.303t \log_{10} \frac{[A]_0}{[A]_t}$  (b)  $K = -\frac{1}{t} \ln \frac{[A]_t}{[A]_0}$   
(c)  $K = \frac{2.303}{t} \log_{10} \frac{[A]_t}{[A]_0}$  (d)  $K = \frac{1}{t} \ln \frac{[A]_t}{[A]_0}$

- Q. 2. Answer any SIX of the following : [12]**
- (i) Define the following terms :  
(a) Enthalpy of fusion  
(b) Enthalpy of atomization
- (ii) Derive van't Hoff general solution equation.
- (iii) Explain impurity defect in stainless steel with diagram.
- (iv) Derive the relation between half life and rate constant for a first order reaction.
- (v) Draw neat and labelled diagram of dry cell.
- (vi) Explain the structure of sulphur dioxide.
- (vii) What is calcination ? Explain it with reactions.
- (viii) Arrange the following reducing agents in the order of increasing strength under

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**M Walker**



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## **Table of Contents H S C Chemistry C Q Question 2015**

1. Understanding the eBook H S C Chemistry C Q Question 2015
  - The Rise of Digital Reading H S C Chemistry C Q Question 2015
  - Advantages of eBooks Over Traditional Books
2. Identifying H S C Chemistry C Q Question 2015
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an H S C Chemistry C Q Question 2015
  - User-Friendly Interface
4. Exploring eBook Recommendations from H S C Chemistry C Q Question 2015
  - Personalized Recommendations
  - H S C Chemistry C Q Question 2015 User Reviews and Ratings
  - H S C Chemistry C Q Question 2015 and Bestseller Lists
5. Accessing H S C Chemistry C Q Question 2015 Free and Paid eBooks
  - H S C Chemistry C Q Question 2015 Public Domain eBooks
  - H S C Chemistry C Q Question 2015 eBook Subscription Services
  - H S C Chemistry C Q Question 2015 Budget-Friendly Options
6. Navigating H S C Chemistry C Q Question 2015 eBook Formats

- ePub, PDF, MOBI, and More
  - H S C Chemistry C Q Question 2015 Compatibility with Devices
  - H S C Chemistry C Q Question 2015 Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of H S C Chemistry C Q Question 2015
    - Highlighting and Note-Taking H S C Chemistry C Q Question 2015
    - Interactive Elements H S C Chemistry C Q Question 2015
  8. Staying Engaged with H S C Chemistry C Q Question 2015
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers H S C Chemistry C Q Question 2015
  9. Balancing eBooks and Physical Books H S C Chemistry C Q Question 2015
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection H S C Chemistry C Q Question 2015
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine H S C Chemistry C Q Question 2015
    - Setting Reading Goals H S C Chemistry C Q Question 2015
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of H S C Chemistry C Q Question 2015
    - Fact-Checking eBook Content of H S C Chemistry C Q Question 2015
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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