

# 5.2 Exercises

Dynamic Solutions available at [BigIdeasMath.com](http://BigIdeasMath.com)

In Exercises 5–8,  $\triangle XYZ \cong \triangle MNL$ . Copy and complete the statement.

5.  $m\angle Y = \underline{\hspace{2cm}}$

6.  $m\angle M = \underline{\hspace{2cm}}$

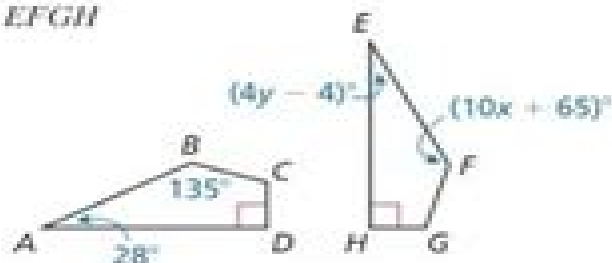
7.  $m\angle Z = \underline{\hspace{2cm}}$

8.  $XY = \underline{\hspace{2cm}}$



In Exercises 9 and 10, find the values of  $x$  and  $y$ .  
(See Example 2.)

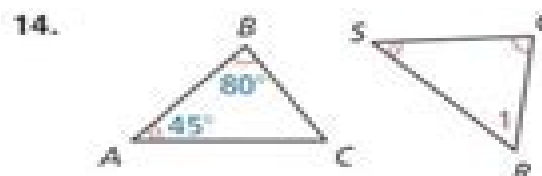
9.  $ABCD \cong EFGH$



10.  $\triangle MNP \cong \triangle TUS$



In Exercises 13 and 14, find  $m\angle 1$ . (See Example 4.)



**ERROR ANALYSIS** In Exercises 17 and 18, describe and correct the error.

17. Given  $\triangle QRS \cong \triangle XYZ$

$\angle S \cong \angle Z$   
 $m\angle S = m\angle Z$   
 $m\angle S = 42^\circ$

**MATHEMATICAL CONNECTIONS** In Exercises 23 and 24, use the given information to write and solve a system of linear equations to find the values of  $x$  and  $y$ .

23.  $\triangle LMN \cong \triangle PQR$ ,  $m\angle L = 40^\circ$ ,  $m\angle M = 90^\circ$ ,  
 $m\angle P = (17x - y)^\circ$ ,  $m\angle R = (2x + 4y)^\circ$

24.  $\triangle STU \cong \triangle XYZ$ ,  $m\angle T = 28^\circ$ ,  $m\angle U = (4x + y)^\circ$ ,  
 $m\angle X = 130^\circ$ ,  $m\angle Y = (8x - 6y)^\circ$

# Geometry Springboard Answer Key

**Tobias Bleicker**



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web we used mole fractions in gen chem 1 to describe gas mixtures where the partial pressure of a gas was its mole fraction time the total pressure  $p_a = x_a p_{\text{total}}$  mole fraction  $x_a = \frac{\text{moles } a}{\text{moles total}}$   $\sum x_i = 1$  mole fraction of substance  $a$  is often denoted by  $x_a$  and

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web 1 in the given mixture the sum of all the mole fractions is equal to one  $x_a + x_b = 1$  2 the multiplication of the mole fraction by 100 gives the mole percentage 3 mole fraction is a unitless and dimensionless expression advantages of molar fraction mole fraction is not dependent on the temperature

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web nov 6 2023  $x_a = \frac{\text{mol } a}{\text{mol } a + \text{mol } b}$  and  $x_b = \frac{\text{mol } b}{\text{mol } a + \text{mol } b}$  note in the given mixture the sum of all the mole fractions present is equal to one  $x_a + x_b = 1$  when the mole fractions are multiplied by 100 they give the mole percentage mole fraction is a unitless and dimensionless expression image will be uploaded soon

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web jan 30 2020 ans the molality of solution is 1 mol kg<sup>-1</sup> or 1 m example 0 211 g of urea  $NH_2CONH_2$  was dissolved in 100 g of water calculate the molarity and molality of the solution given n 14 h 1 c 12 o 16 given mass of solute urea 11 11 g mass of solvent water 100 g 0 1 kg to find molarity of solution solution

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web you may work out the mole fractions on your own problem 5 what is the mass of a sample of a 0 449 molal kbr that contains 2 92 kg of water 1 molality moles solute divided by kilograms solute 0 449 mol kg x 2 92 kg x 1 31108 mol of kbr 2 moles times molar mass equals grams 1 31108 mol 119 0023 g mol 156 g kbr

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web calculate the mass percent mole fraction molality and molarity of hexane  $C_6H_{14}$  660 g cm<sup>-3</sup> when it is mixed with heptane  $C_7H_{16}$  684 g cm<sup>-3</sup> in a 35 0 55 0 ml ratio assume that the volumes are additive

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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web jan 8 2022 mol % x 100 the sum of all of the mole percents of a mixture equals 100 mole fraction properties and advantages the mole fraction offers advantages over some of the other units of concentration unlike molarity mole fraction is

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