



# CHEM FAX!

FLINN  
LABORATORY EQUIPMENT

## Preparing and Diluting Solutions Student Laboratory Kit

### Introduction

Solutions are homogeneous mixtures. They have the same composition and properties throughout. Solutions are formed by dissolving a solute in a solvent.

### Concepts

- 1. Solubility
- 2. Concentration

### Background

#### Preparing Solutions

The amount of solute that can be dissolved in a solvent is called the solubility. The amount of solute that is dissolved in a solvent is called the concentration. The concentration of a solution can be expressed in many ways.

One way to express concentration is by mass. The mass of solute divided by the mass of the solution is called the mass percent. The mass percent of a solution is the mass of solute divided by the mass of the solution, multiplied by 100.

Another way to express concentration is by volume. The volume of solute divided by the volume of the solution is called the volume percent. The volume percent of a solution is the volume of solute divided by the volume of the solution, multiplied by 100.

Another way to express concentration is by molarity. The number of moles of solute divided by the volume of the solution in liters is called the molarity. The molarity of a solution is the number of moles of solute divided by the volume of the solution in liters.

SAFETY: Wear safety glasses and gloves. Do not drink or eat from the laboratory.



# Flinn Scientific Preparing And Diluting Answers

**G Thomas**



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