

Substance	Lewis Structure	Molar Mass	Intermolecular Attractive Force
Methanol CH_3OH $12 + 3(1) + 16 + 1 = 32$	$\begin{array}{c} \text{H} \\ \\ \text{H} - \text{C} - \ddot{\text{O}} - \text{H} \\ \\ \text{H} \end{array}$	32.042	Hydrogen Bonds
Ethanol $\text{CH}_3\text{CH}_2\text{OH}$ $12 + 5(1) + 16 + 1 = 46$	$\begin{array}{c} \text{H} \quad \text{H} \quad \\ \quad \quad \\ \text{H} - \text{C} - \text{C} - \ddot{\text{O}} - \text{H} \\ \quad \\ \text{H} \quad \text{H} \end{array}$	46.068	Hydrogen Bonds
n-Butanol $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$	$\begin{array}{c} \text{H} \quad \text{H} \quad \text{H} \quad \text{H} \quad \text{H} \\ \quad \quad \quad \quad \\ \text{H} - \text{C} - \text{C} - \text{C} - \text{C} - \ddot{\text{O}} - \text{H} \\ \quad \quad \quad \\ \text{H} \quad \text{H} \quad \text{H} \quad \text{H} \end{array}$	74.122	Hydrogen Bonds
Glycerin $\text{C}_3\text{H}_8(\text{OH})_3$ $12 + 5(1) + 18(16) = 92$	$\begin{array}{c} \text{H} \quad \text{:}\ddot{\text{O}}\text{:} - \text{H} \quad \text{H} \\ \quad \quad \\ \text{H} - \text{C} - \text{C} - \text{C} - \text{H} \\ \quad \quad \\ \text{:}\ddot{\text{O}}\text{:} - \text{H} \quad \text{H} \quad \text{:}\ddot{\text{O}}\text{:} - \text{H} \end{array}$	92.093	Hydrogen Bonds
Water H_2O	$\begin{array}{c} \text{H} \\ \\ \text{H} - \ddot{\text{O}} - \text{H} \\ \\ \text{H} \end{array}$	18.015	Hydrogen Bonds

Exploring Intermolecular Forces Lab Answers

AW Chickering



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