



Activity B:

Voltage gradients

Get the Gizmo ready:

- Set the **Wavelength** to 300 nm, the **Photon flux** to 10 γ /ms, and the **Voltage** to 0.0 volts.
- Turn on **Show voltage gradient**.



Introduction: The electrons that are freed from the surface of the metal have a specific amount of kinetic energy. Faster electrons have greater energies than slower ones. The energy of emitted electrons is measured by setting up an electrical field that opposes their motion. The **voltage** of the field is a measure of its strength.

Goal: Use a voltage gradient to measure the energy of emitted electrons.

1. **Observe:** Check that **Potassium** is selected. Click **Flash the light** and observe the emitted electrons. Increase the **Voltage** to 1.5 volts, and click **Flash the light** again.

How does the electrical field affect the motion of the emitted electrons?

The electrons are caused to slow down as they move through the tube because of the electrical field.

2. **Measure:** The energy of an emitted electron is measured in **electron volts** (eV). An electron with an energy of 1 eV can overcome an electrical field of 1 volt. In the Gizmo, increase the voltage until you find the highest voltage that still allows the electrons to reach the light bulb.

What is this value?

1.8eV

This is equal to the energy of the emitted electrons in eV.

3. **Gather data:** With the **Wavelength** set to 300 nm, measure the energy of emitted electrons for potassium, calcium, and uranium. Then measure the same values with wavelengths of 250 nm and 200 nm to complete the table.

Element	Energy of emitted electrons (eV)		
	300 nm	250 nm	200 nm
Potassium	1.8eV	2.6eV	3.8eV
Calcium	1.2eV	2eV	3.2eV
Uranium	0.5eV	1.3eV	2.5eV

4. **Analyze:** What patterns do you notice in your data?

The wavelength of the light is decreased as the energy of released electrons increases.

5. **Infer:** Based on your data, which element is hardest to extract electrons from?

Uranium

Explain

The electrons emitted from uranium has the least energy at each wavelength

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