

4. **Analyze:** How did the number of solute particles affect the colligative properties of water?

The boiling point increased and the freezing point decreased as the number of particles increased.

5. **Find a pattern:** The chemical formulas of the three ionic solutes are NaCl (sodium chloride), CaCl₂ (calcium chloride), and KCl (potassium chloride).

- A. How does the chemical formula relate to the number of particles that result from the dissociation of a molecule?

The compound's particle count multiplied by the number of particles in solution the variety of the elements.

- B. Aluminum chloride has the formula AlCl₃. How many particles would a molecule of aluminum chloride dissociate into?

8

- C. Based on the data you have collected so far, predict the boiling point and freezing point of a 1.0-mol/kg solution of aluminum chloride:

16

6. **Check:** Select the STUDY CONCENTRATION tab. You can model a 1.0-mol/kg solution of aluminum chloride by using a 2.0-mol/kg solution of sodium chloride. Use the Gizmo to find the boiling point and freezing point of a 1.0-mol/kg aluminum chloride solution.

Solute	Number of particles	Vapor pressure at 25 °C	Boiling point	Freezing point
Aluminum chloride	8	1.03	102	-7

7. **Think and discuss:** Colligative properties depend only on the number of particles, not the identity of the solute. What evidence could you use to demonstrate this point?

The same events do not occur with the same identities when the same number of particles have the same colligative qualities.

Explore Learning Colligative Properties Answers

RS Peters



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