

12-17-07

P 178 Name an advantage to each model: the ball-and-stick model and the space-filling model.

A ball and stick model shows bond length, and a space-filling model shows relative atom sizes.

P 180 How is the hardness of minerals explained by minerals' chemical structure?

The hardness of minerals is explained by the fact that their chemical structure consists of rigid networks of bonded atoms.

Section 1 Review page 182

1. Classify the following substances as mixtures or compounds

- air: mixture
- CO: compound
- SnF₂: compound
- pure water: compound

2. Draw a ball and stick model of a boron trifluoride, BF₃, molecule. In this molecule, a boron atom is attached to three fluorine atoms. Each F-B-F bond angle is 120 degrees, and all B-F bonds are the same length.

3. Explain why silver iodide, AgI, a compound used in photography, has a much higher melting point than vanillin, C₈H₈O₃, a sweet-smelling compound used in flavoring.

Silver iodide has a much higher attraction between particles, a structure of positive of negative ions. Vanillin consists of molecules.

4. Explain why glass, which is made of mostly SiO₂, is often used to make cookware.

SiO₂ has a high melting point. It does not melt when heated to high cooking temperatures.

5. Predict which molecules have a greater attraction for each other: C₃H₈O molecules in liquid rubbing alcohol or CH₄ molecules in methane gas.

The C₃H₈O have a greater attraction than the CH₄ because they are closer together and they are moving more slowly than those in a gas.

6. A pictometer (pm) is equal to 1×10^{-12} m. The O-H bond lengths in water are 95.8 pm, while S-H bond lengths in dihydrogen sulfide are 135 pm. Why are S-H bonds longer than O-H bonds?

The sulfur atoms are larger than oxygen atoms by one electron energy level.

P 184: What holds two ionic bonded atoms together?

In an ionic bond, ions are held together by the attraction between their opposite charges.

P 186: What holds two covalently bonded atoms together?

Holt Physical Science Chapter 6 Review Answers

Padhraic Smyth



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