

AP QUESTION - 1995

- (a) mole fraction CO
assume 1.0 L



$$Y_{\text{CO}} = \frac{.55 \text{ mol CO}}{1.6 \text{ mol total}} = .34375 = \boxed{.34}$$

$$(b) K_c = \frac{[\text{H}_2\text{O}][\text{CO}]}{[\text{H}_2][\text{CO}_2]} = \frac{(.55 \text{ M})(.55 \text{ M})}{(.20 \text{ M})(.30 \text{ M})} = \boxed{5.04} \quad (\text{no unit})$$

$$(c) K_p = K_c (RT)^{\Delta n} \quad \Delta n = 0 \quad \therefore K_p = K_c = \boxed{5.04}$$

(d)

$$\text{H}_2 + \text{CO}_2 \rightleftharpoons \text{H}_2\text{O} + \text{CO}$$

i	.20 M	.30 M	.55 M	.55 M
Δ	+ .165 M	+ .165 M	- .165 M	- .165 M
eq	.365	.465	.385	.385

$$30\% \text{ of } .55 = .165$$

$$K_c = \frac{[\text{H}_2\text{O}][\text{CO}]}{[\text{H}_2][\text{CO}_2]} = \frac{(.385)(.385)}{(.365)(.465)} = \boxed{.873}$$

(e)

i	.50 / $\sqrt{2}$	.50 / $\sqrt{2}$	0 M	0 M
Δ	-x	-x	+x	+x
eq	.167 - x	.167 - x	x	x

$$\frac{x^2}{(.167 - x)^2} = 5.04$$

$$\frac{x}{.167 - x} = \sqrt{5.04}$$

$$\frac{x}{.167 - x} = 2.24$$

$$\boxed{x = .116 \text{ M CO}}$$

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