



QUESTION 11 (worth 10%)

Time taken:

Score: 10/10 (100%)

10/10

10/10

10/10

10/10

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10/10

Your Result: correct for

10/10

10/10

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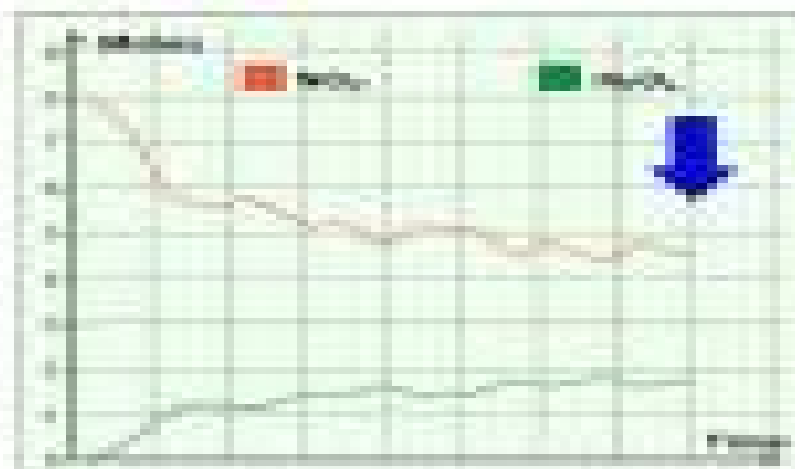
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score 100% (10/10)

Question 11 (worth 10%)

A closed flask containing hydrogen gas at high pressure (100 atm) and some hydrogen molecules (molecules) at low pressure (1 atm). The flask is then cooled. The temperature of the flask is then lowered. What is the effect on the rate of the reaction?



- A. The rate of reaction will be lower than the rate of reaction before.
- B. The rate of reaction will be higher than the rate of reaction before.
- C. The rate of reaction will be the same as the rate of reaction before.
- D. The rate of reaction will be higher than the rate of reaction before.

Correct Answer: A. The rate of reaction will be lower than the rate of reaction before.

Explanation: The rate of reaction is determined by the activation energy (E_a) and the concentration of the reactants. In this case, the concentration of the reactants is constant, but the activation energy is increased when the temperature is lowered. This results in a lower rate of reaction.

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Minjie Lin

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