



Name: _____

Date: _____

Student Exploration: Calorimetry Lab

Vocabulary: calorie, calorimeter, joule, specific heat capacity

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. The Latin word *calor* means “heat,” and *meter* comes from the Greek word meaning “to measure.” What do you think a **calorimeter** does? to measure heat

2. Where have you heard the word **calorie** before? What do you think a calorie is? _____
when looking at the nutritional facts on the back of a food container

Gizmo Warm-up

A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up.

Calorimeters can be used to find a substance’s **specific heat capacity**. You will use the *Calorimetry Lab* Gizmo to determine the specific heat capacities of various substances.



1. On the **SIMULATION** pane, select **Copper**. Use the slider to set its **Mass** to 200 g. Set the **Water mass** to 200 g. Check that the **Water temp** is set to 30.0 °C and the copper’s **Temp** is 90 °C. Select the **GRAPH** tab, and click **Play** (▶).
 - A. What was the **Final temperature** of the copper and the water? 34.96
 - B. How much did the temperature of the copper change? 55.04
 - C. How much did the temperature of the water change? 4.96
2. Specific heat capacity can be described as a substance’s resistance to temperature changes. Which substance has a greater specific heat capacity, copper or water? Explain.
water, because its temperature only changed by 4.96 compared to 55.04



Explore Learning Calorimetry Gizmo Answers

SJ Ball



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Table of Contents Explore Learning Calorimetry Gizmo Answers

1. Understanding the eBook Explore Learning Calorimetry Gizmo Answers
 - The Rise of Digital Reading Explore Learning Calorimetry Gizmo Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Explore Learning Calorimetry Gizmo Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Explore Learning Calorimetry Gizmo Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Explore Learning Calorimetry Gizmo Answers
 - Personalized Recommendations
 - Explore Learning Calorimetry Gizmo Answers User Reviews and Ratings
 - Explore Learning Calorimetry Gizmo Answers and Bestseller Lists

5. Accessing Explore Learning Calorimetry Gizmo Answers Free and Paid eBooks
 - Explore Learning Calorimetry Gizmo Answers Public Domain eBooks
 - Explore Learning Calorimetry Gizmo Answers eBook Subscription Services
 - Explore Learning Calorimetry Gizmo Answers Budget-Friendly Options
6. Navigating Explore Learning Calorimetry Gizmo Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Explore Learning Calorimetry Gizmo Answers Compatibility with Devices
 - Explore Learning Calorimetry Gizmo Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Explore Learning Calorimetry Gizmo Answers
 - Highlighting and Note-Taking Explore Learning Calorimetry Gizmo Answers
 - Interactive Elements Explore Learning Calorimetry Gizmo Answers
8. Staying Engaged with Explore Learning Calorimetry Gizmo Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Explore Learning Calorimetry Gizmo Answers
9. Balancing eBooks and Physical Books Explore Learning Calorimetry Gizmo Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Explore Learning Calorimetry Gizmo Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Explore Learning Calorimetry Gizmo Answers
 - Setting Reading Goals Explore Learning Calorimetry Gizmo Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Explore Learning Calorimetry Gizmo Answers
 - Fact-Checking eBook Content of Explore Learning Calorimetry Gizmo Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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